

CHEMISTRY Objective Final Exam Review 10-11

There are 50 questions on the test. The chapters are listed below. Each of the letters represents a concept/idea that you should know or look up.

1. CH 6: Periodic trends, where do they increase & decrease
 - a. Atomic radius & Ionic radius
 - b. Electronegativity
 - c. Electron affinity
 - d. Ionization energy
 - e. Reactivity of metals and of non-metals
2. CH 7: Look up properties of elements per group and the names of each group.
 - a. Group 1 very reactive in water; group 7 most reactive non-metals etc.
 - b. Fluorine most reactive in periodic table
 - c. Group 1 will lose e- to become +1
 - d. Elements in same group have same chem properties.
3. CH 8: Ionic compounds and bonds
 - a. Why do atoms bond? How many electrons make them stable?
 - b. How to name ionic compounds, go from name to formula
 - c. What elements make up ionic compounds?
 - d. Properties of ionic compounds
 - e. What is a crystal lattice and the arrangement.
 - f. What is metallic bonding?
4. CH 9: Covalent compounds
 - a. What elements make up covalent compounds?
 - b. Compare ionic vs covalent bonds. An e-neg difference of 1.7 is dividing factor.
 - c. What are pi bonds and sigma bonds?
 - d. Bond length Vs. Bond strength, and Bond dissociation energy.
 - e. What are molecules?
 - f. What are the common diatomic molecules?
 - g. Polar molecules, partial charges, what are they?
5. CH 10: Identify the types of chemical reactions
 - a. Be able to write the products for a double replacement and single replacement reactions.
 - b. Balancing equations
 - c. Writing equations from word names.
6. CH 11: The mole
 - a. Determine the molar mass of a compound
 - b. Calculate the amt of particles, ions, atoms, formula units, molecules in a compound.
 - c. Converting mole \rightarrow mole, mole \rightarrow grams, mole \rightarrow particles and visa versa
 - d. Converting grams \rightarrow mole, grams \rightarrow grams, grams \rightarrow particles visa vera
7. CH 12:
 - a. Stoichiometry conversions using balanced chemical equation
 - b. Mole \rightarrow mole, mole \rightarrow grams, grams \rightarrow grams

Objective Final Review For Chemistry

R Barnett



Objective Final Review For Chemistry:

OAT Prep Plus 2019-2020 Kaplan Test Prep, 2018-12-04 Kaplan's OAT Prep Plus 2019 2020 provides the test taking strategies realistic practice and expert guidance you need to get the OAT results you want Our comprehensive updated subject review reflects recent changes to the blueprint of the exam question types and test interface You'll get two full length practice OATs and expert tips to help you face Test Day with confidence The Best Review Two updated full length online practice exams for test like practice Study planning guidance More than 600 practice questions for every subject with detailed answers and explanations Full color study sheets for high yield review on the go A guide to the current OAT Blueprint so you know exactly what to expect on Test Day Comprehensive review of all of the content covered on the OAT Expert Guidance Our books and practice questions are written by veteran teachers who know students every explanation is written to help you learn Kaplan's experts ensure our practice questions and study materials are true to the test We invented test prep Kaplan www.kaptest.com has been helping students for 80 years and our proven strategies have helped legions of students achieve their dreams The previous edition of this book was titled OAT 2017 2018 Strategies Practice Review

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