

# Part A

Answer all questions in this part.

Directions (1–30): For each statement or question, record on your separate answer sheet the number of the word or expression that, of those given, best completes the statement or answers the question. Some questions may require the use of the 2011 Edition Reference Tables for Physical Setting/Chemistry.

1 Which statement describes the charge of an electron and the charge of a proton?

- (1) An electron and a proton both have a charge of +1.
- (2) An electron and a proton both have a charge of -1.
- (3) An electron has a charge of +1, and a proton has a charge of -1.
- (4) An electron has a charge of -1, and a proton has a charge of +1.

2 Which subatomic particles are found in the nucleus of an atom of beryllium?

- (1) electrons and protons
- (2) electrons and positrons
- (3) neutrons and protons
- (4) neutrons and electrons

3 The elements in Period 4 on the Periodic Table are arranged in order of increasing

- (1) atomic radius
- (2) atomic number
- (3) number of valence electrons
- (4) number of occupied shells of electrons

4 Which phrase describes two forms of solid carbon, diamond and graphite, at STP?

- (1) the same crystal structure and the same properties
- (2) the same crystal structure and different properties
- (3) different crystal structures and the same properties
- (4) different crystal structures and different properties

5 Which element has six valence electrons in each of its atoms in the ground state?

- (1) Se
- (2) As
- (3) Kr
- (4) Ga

6 What is the chemical name for  $\text{H}_2\text{SO}_3(\text{aq})$ ?

- (1) sulfuric acid
- (2) sulfurous acid
- (3) hydrosulfuric acid
- (4) hydrosulfurous acid

7 Which substance is most soluble in water?

- (1)  $(\text{NH}_4)_3\text{PO}_4$
- (2)  $\text{Cu}(\text{OH})_2$
- (3)  $\text{Ag}_2\text{SO}_4$
- (4)  $\text{CaCO}_3$

8 Which type of bonding is present in a sample of an element that is malleable?

- (1) ionic
- (2) metallic
- (3) nonpolar covalent
- (4) polar covalent

9 Which atom has the greatest attraction for the electrons in a chemical bond?

- (1) hydrogen
- (2) oxygen
- (3) silicon
- (4) sulfur

10 Which type of reaction involves the transfer of electrons?

- (1) alpha decay
- (2) double replacement
- (3) neutralization
- (4) oxidation-reduction

11 A 10.0-gram sample of nitrogen is at STP. Which property will increase when the sample is cooled to 72 K at standard pressure?

- (1) mass
- (2) volume
- (3) density
- (4) temperature

12 Which element is a gas at STP?

- (1) sulfur
- (2) xenon
- (3) potassium
- (4) phosphorus

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**Rachel S Tattersall**



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