# Chemistry Unit 8 Worksheet 4 Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

 Calculate the number of moles of potassium chlorate, KClO<sub>3</sub> (s), that must decompose to produce potassium chloride, KCl (s), and 1.8 moles of oxygen gas.

Equation:	2 KClO <sub>3</sub> →	2 KCI +	3 O2
Before:	1.2	•	0
Change:	-1.2	+1.2	+1.8
After:	0	1.2	1.8

1.2 moles KClO<sub>3</sub> needed

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

Equation:	Mig -	2 HCl →	MgCl <sub>2</sub> +	FE 2
Before:	0.100	7	0	0
Change:	-0.100	-0.200 +	0.160 +	0.100
Afters	0	.2	0.100	0.100

0.200 moles HCI needed

0.73 g x 1 mole = .0040 mole 183 g

3. Ethane, C<sub>2</sub>H<sub>6</sub> reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

Equation:	2 C2H6	+ 7 O <sub>2</sub>	→ 4 CO;	+ 6 H <sub>2</sub> O
Before:	2.20	26.26	0	0
Change:	-2.20	- 7.70	+ 4,40	+ 6.60
After:	0	2.5	4.40	6.60

O<sub>2</sub> needed: 7.70 mole x 32.0 g = 246 g O<sub>2</sub> required I mole

 Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:

Ni(NO <sub>3</sub> ) <sub>2</sub> + 2 NaOH> Ni(OH) <sub>2</sub> + 2 NaNO <sub>3</sub>				0,	
Equation:	$Ni(NO_3)_2 +$	2 NaOH	→ Ni(OH) <sub>2</sub>	+ 2 NaNO3	
Before:	0.0040	XS	0	0	

The state of the s		The Property of the Section 1972	The second second
0.0040	NS	0	0
- 0.0040	- 0.0080	+0.0040	+0.0080
0	NS	0.0040	0.0080
	0.0040	0.0040 XS	0.0040 XS 0 - 0.0040 - 0.0080 +0.0040

NaNO<sub>3</sub> produced: 0.0080 mole x <u>85.0 g</u> = 0.68 g 1 mole

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**Lauren Gardner** 

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