

C1.1 Atomic structure

Question number	Answer	Marks	Guidance
1 a	metals: Ba, V, Hg, K, U; non-metals: P, Kr	1 1	
1 b	mixture: atoms not chemically bonded	1	
2	compound: atoms chemically bonded	1	
3	element: all atoms same type; compound: more than one type of atom chemically combined	1 1	
4	tiny central nucleus surrounded by orbiting electrons	1 1	
5 a	radium	1	
5 b	zinc	1	
5 c	plutonium	1	
5 d	caesium	1	
6	twice as many oxygen atoms as carbon atoms / carbon and oxygen atoms in carbon dioxide bond to each other in the ratio 1 : 2 (carbon : oxygen)	2	

C1.2 Chemical equations

Question number	Answer	Marks	Guidance
1 a	No atoms are created or destroyed in a chemical reaction so equal numbers of each type of atom must appear on both sides of the equation.	1 1	
1 b	$\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$	1	
2 a	17.6 g	1	
2 b	$\text{MgCO}_3(\text{s}) \rightarrow \text{MgO}(\text{s}) + \text{CO}_2(\text{g})$	1	
3 a	$2\text{KNO}_3 \rightarrow 2\text{KNO}_2 + \text{O}_2$	1	
3 b	$4\text{Li} + \text{O}_2 \rightarrow 2\text{Li}_2\text{O}$	1	
3 c	$4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$	1	
3 d	$\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$	1	
4	$2\text{Na}(\text{s}) + 2\text{H}_2\text{O}(\text{l}) \rightarrow 2\text{NaOH}(\text{aq}) + \text{H}_2(\text{g})$	2	

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