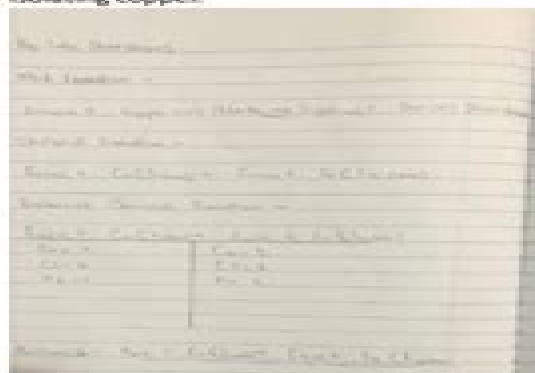


Limiting Reagents and % Yield

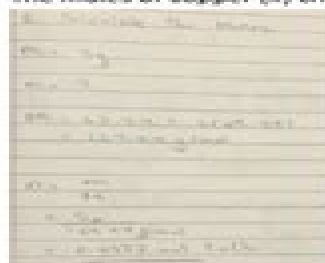
The percentage yield of a chemical reaction is determined by numerous factors: the nature of the reaction itself, the conditions under which the reaction was carried out and the nature of the reactants used. In this investigation, you will determine the limiting/excess reagents and the percentage yield of the reaction between iron and a solution of copper (II) chloride.

Prelab Questions:

- Write the balanced chemical equation for the reaction between iron and copper (II) chloride. Remember to use your activity series to check the products.
 - The balanced equation is $\text{Fe(s)} + \text{CuCl}_2\text{(aq)} \rightarrow \text{Cu(s)} + \text{FeCl}_2\text{(aq)}$. The reaction is a single displacement isolating copper.



- Calculate the moles of copper (II) chloride in the solution if you use 5 g of copper (II) chloride.
 - The moles of copper (II) chloride in the solution, if I use 5g of copper (II) chloride, is 0.037 mol CuCl_2 .



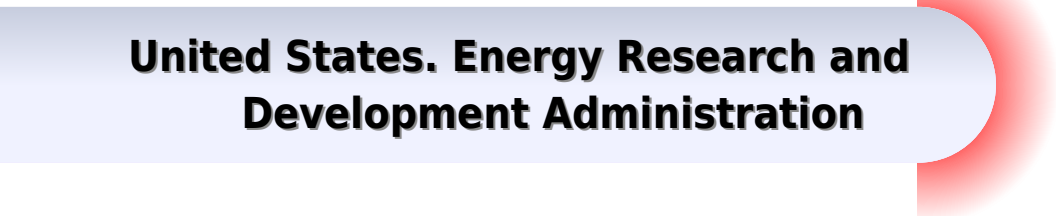
- Determine the limiting reagent if you use 1 g of iron.
 - The limiting reagent, if I use 1g of iron, is Fe(s) . It is 0.018mol because when comparing the numbers by the

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web description you may be required to comply the code if you intend to design and construct buildings and other structures at an airport site the building code of australia bca

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