

Periodicity - Electronegativity

Name: **Answer Key**

1. Fill in the blanks below, and then draw arrows on to the periodic table indicating the direction of the electronegativity trend.

Generally, electronegativity

increases

across periods and

increases

up groups.



2. List the elements below in order of increasing electronegativity. (Smallest to largest).

Ba, Rb, Ca, Sr, Mg

Rb, Ba, Ca, Mg, Sr

Zn, Ni, Cu, Sc, Ti

Sc, Ti, Ni, Cu, Zn

F, B, O, Rb, Fr

Fr, Rb, B, O, F

Nb, Sn, Rb, Fr, I

Fr, Rb, Nb, Sn, I

F, Li, Ba, K, Na

K, Na, Li, Ba, F

P, Se, Ar, Mg, Ca

Ca, Mg, Se, P, Ar

3. Are the elements below in order of decreasing electronegativity? If yes, write "yes". If not, write them in the correct order.

Ca, K, Se, Br

Br, Se, Ca, K

I, Br, Cl, F

F, Cl, Br, I

O, B, Li, Na

Yes

Y, Sr, Rb, Fr

Yes

At, Sc, Se, Ba

Yes

As, Pb, Te, Si

Si, Te, Pb, As

4. Briefly explain why electronegativity changes up groups.

Generally, electronegativity increases as you move up in a group. This is because the atomic number decreases, and as a result, there is less distance between the nucleus and valence electrons. In other words, there is a smaller atomic radius.

5. Briefly explain why electronegativity changes across periods.

As you move from left to right, across a period, the nuclear charge of elements increases faster than electron shielding. This results in an increased attraction to electrons.

Periodicity Worksheet Answer Guide

IM Harris



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Table of Contents Periodicity Worksheet Answer Guide

1. Understanding the eBook Periodicity Worksheet Answer Guide
 - The Rise of Digital Reading Periodicity Worksheet Answer Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodicity Worksheet Answer Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Periodicity Worksheet Answer Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodicity Worksheet Answer Guide
 - Personalized Recommendations
 - Periodicity Worksheet Answer Guide User Reviews and Ratings
 - Periodicity Worksheet Answer Guide and Bestseller Lists
5. Accessing Periodicity Worksheet Answer Guide Free and Paid eBooks
 - Periodicity Worksheet Answer Guide Public Domain eBooks
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 - Periodicity Worksheet Answer Guide Budget-Friendly Options
6. Navigating Periodicity Worksheet Answer Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodicity Worksheet Answer Guide Compatibility with Devices
 - Periodicity Worksheet Answer Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodicity Worksheet Answer Guide
 - Highlighting and Note-Taking Periodicity Worksheet Answer Guide
 - Interactive Elements Periodicity Worksheet Answer Guide

8. Staying Engaged with Periodicity Worksheet Answer Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodicity Worksheet Answer Guide
9. Balancing eBooks and Physical Books Periodicity Worksheet Answer Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodicity Worksheet Answer Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodicity Worksheet Answer Guide
 - Setting Reading Goals Periodicity Worksheet Answer Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodicity Worksheet Answer Guide
 - Fact-Checking eBook Content of Periodicity Worksheet Answer Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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