

Periodic Trends Worksheet

Directions: Use your notes to answer the following questions.

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to iodine. This results in smaller size of Fluorine than iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- | | |
|--|------------------|
| a. atomic radius (excluding noble gases) | Decreases |
| b. first ionization energy | Increases |
| c. electronegativity | Increases |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **increases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (ionization Energy) increases.

Periodic Trends And Atomic Properties Pogil Worksheet

Answer Key

Xiaolong Qi



Periodic Trends And Atomic Properties Pogil Worksheet Answer Key:

, *Science Spectrum* Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2004 **Chemistry and Our Universe** , Return to the periodic table introduced in Lecture 1 to practice predicting properties of elements based on their electronic structure Then witness what happens when three different alkali metals react with water Theory forecasts a pronounced difference in the result Is there *Atomic Properties of the Elements* , 2001 *The periodic table. 4th ed* ,

The book delves into Periodic Trends And Atomic Properties Pogil Worksheet Answer Key. Periodic Trends And Atomic Properties Pogil Worksheet Answer Key is a vital topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Periodic Trends And Atomic Properties Pogil Worksheet Answer Key, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:
 - Chapter 1: Introduction to Periodic Trends And Atomic Properties Pogil Worksheet Answer Key
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 - Chapter 3: Periodic Trends And Atomic Properties Pogil Worksheet Answer Key in Everyday Life
 - Chapter 4: Periodic Trends And Atomic Properties Pogil Worksheet Answer Key in Specific Contexts
 - Chapter 5: Conclusion
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 6. In chapter 5, the author will draw a conclusion about Periodic Trends And Atomic Properties Pogil Worksheet Answer Key. This chapter will summarize the key points that have been discussed throughout the book.
- The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Periodic Trends And Atomic Properties Pogil Worksheet Answer Key.

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