

CHAPTER 6

STUDY GUIDE

The Periodic Table and Periodic Law

Section 6.1 Development of the Modern Periodic Table

In your textbook, read about the history of the periodic table's development.

Use each of the terms below just once to complete the passage.

~~octaves~~~~atomic mass~~~~atomic number~~~~size~~~~elements~~~~properties~~~~Henry Moseley~~~~eight~~

protons

periodic law

Dmitri Mendeleev

accepted

The table below was developed by John Newlands and is based on a relationship called the law of (1) properties. According to this law, the properties of the elements repeated every (2) eight elements. Thus, for example, element two and element (3) nine have similar properties. The law of octaves did not work for all the known elements and was not generally (4) accepted.

1	2	3	4	5	6	7
H	Li	G	Be	C	N	O
8	9	10	11	12	13	15
F	Na	Mg	Al	Si	P	S

The first periodic table is mostly credited to (5) Mendeleev. In his table, the elements were arranged according to increasing (6) atomic mass. One important result of this table was that the existence and properties of undiscovered (7) elements could be predicted.

The element in the modern periodic table are arranged according to increasing (8) atomic number, as a result of the work of (9) Moseley. This arrangement is based on number of (10) protons in the nucleus of an atom of the element. The modern form of the periodic table results in the (11) periodic law which states that when elements are arranged according to increasing atomic number, there is a periodic repetition of their chemical and physical (12) properties.

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Table of Contents Periodic Law Study Guide Answers

1. Understanding the eBook Periodic Law Study Guide Answers
 - The Rise of Digital Reading Periodic Law Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Law Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Law Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Law Study Guide Answers
 - Personalized Recommendations
 - Periodic Law Study Guide Answers User Reviews and Ratings
 - Periodic Law Study Guide Answers and Bestseller Lists
5. Accessing Periodic Law Study Guide Answers Free and Paid eBooks

- Periodic Law Study Guide Answers Public Domain eBooks
- Periodic Law Study Guide Answers eBook Subscription Services
- Periodic Law Study Guide Answers Budget-Friendly Options
- 6. Navigating Periodic Law Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Law Study Guide Answers Compatibility with Devices
 - Periodic Law Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Law Study Guide Answers
 - Highlighting and Note-Taking Periodic Law Study Guide Answers
 - Interactive Elements Periodic Law Study Guide Answers
- 8. Staying Engaged with Periodic Law Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Law Study Guide Answers
- 9. Balancing eBooks and Physical Books Periodic Law Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Law Study Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Law Study Guide Answers
 - Setting Reading Goals Periodic Law Study Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Law Study Guide Answers
 - Fact-Checking eBook Content of Periodic Law Study Guide Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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