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**Michal Rosen-Zvi**



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web answer  $\text{Na}^+ \text{aq}$  and  $\text{SO}_4^{2-} \text{aq}$  identify the chemical equation the spectator ions  $\text{KOH} \text{aq}$   $\text{HNO}_3$   $\text{KNO}_3 \text{aq}$   $\text{H}_2\text{O}$   $\text{K}^+ \text{aq}$   $\text{OH}^- \text{aq}$   $\text{H}^+ \text{aq}$   $\text{NO}_3^- \text{aq}$   $\text{K}^+ \text{aq}$   $\text{NO}_3^- \text{aq}$   $\text{H}_2\text{O}$  l answer  $\text{K}^+ \text{aq}$  and  $\text{NO}_3^- \text{aq}$  identify the chemical equation the spectator ions  $\text{Pb}^{2+} \text{aq}$   $\text{NO}_3^- \text{aq}$   $\text{PbI}_2 \text{s}$   $\text{KNO}_3 \text{aq}$

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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction  $2\text{Ag}^+ \text{aq}$   $\text{Cr}_2\text{O}_7^{2-} \text{aq}$   $\text{H}_2\text{O}$   $\text{Cr}^{3+} \text{aq}$   $\text{Ag}_2\text{Cr}_2\text{O}_7 \text{s}$  label 4 2 3

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web jan 29 2020 the ionic form of this reaction is  $2\text{Na}^+ \text{aq}$   $2\text{Cl}^- \text{aq}$   $\text{Cu}^{2+} \text{aq}$   $\text{SO}_4^{2-} \text{aq}$   $2\text{Na}^+ \text{aq}$   $\text{SO}_4^{2-} \text{aq}$   $\text{CuCl}_2 \text{s}$  the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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 net ionic equation  $\text{H}^+ \text{aq} + \text{HCO}_3^- \text{aq} \rightarrow \text{H}_2\text{O} \text{l} + \text{CO}_2 \text{g}$  here is a video on spectator ions  
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web the six  $\text{Na}^+$  ions and the six  $\text{Cl}^-$  ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation  $\text{Ba}^{2+} \text{aq} + 2\text{PO}_4^{3-} \text{aq} \rightarrow \text{Ba}_3(\text{PO}_4)_2 \text{s}$

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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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 $\text{AgNO}_3 \text{aq} + \text{KCl} \text{aq} \rightarrow \text{AgCl} \text{s} + \text{KNO}_3 \text{aq}$  total ionic  $\text{Ag}^+ \text{aq} + \text{NO}_3^- \text{aq} + \text{K}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{AgCl} \text{s} + \text{K}^+ \text{aq} + \text{NO}_3^- \text{aq}$  net ionic  $\text{Ag}^+ \text{aq} + \text{Cl}^- \text{aq} \rightarrow \text{AgCl} \text{s}$  2

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web video answer you know spectator ions are the ions present on the both sides of the equation indicate cancels out in the total ionic equations so you know in a case chemical reactions are described with chemical equations which symbolically present the starting materials end products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water  $\text{H}^+ \text{aq} + \text{OH}^- \text{aq} \rightarrow \text{H}_2\text{O} \text{l}$  the spectator ions in  $\text{KOH} \text{aq}$  and  $\text{HNO}_3 \text{aq}$  are  $\text{K}^+ \text{aq}$  and  $\text{NO}_3^- \text{aq}$  how to identify and list the

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