

# Periodic and Groups: Valence Electrons: Mastery

## Periodic and Groups

Elements that have the same number of valence electrons have similar properties, so we put them in groups. The the periodic table groups + columns that look like this. They will see family groups 1A-18A. Hydrogen is in the same period as Helium.

Elements that have electrons in the outer shell have similar properties in the Periodic Table. Periods go left to right. Elements in the same vertical columns

Period	Group		Group										18A	
	1A	2A	3A	4A	5A	6A	7A	8A	9A	10A	11A	12A	13A	14A
1	H	He											Ne	
2	Li	Be	B	C	N	O	F	Ne					Ar	
3	Na	Mg	Al	Si	P	S	Cl	Ar					Kr	
4	K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge

Thinking like the game "Monopoly", answer that an element is in period 3 and group 1.

What element is in Period 2, and Group 16A? Period 2 means 2 columns that is Column 16A. The element is oxygen.

What element is in Group 11A and Period 3?

What element is in Group 2A and Period 2?

What group number is in Column 17?

What group number is in Column 14?

## Valence Electrons

Valence Electrons are the most important electrons in an atom. Each group contains has the same number of valence electrons. Valence electrons are the electrons that are involved in chemical bonding.

Hydrogen 1



1 valence electron

Helium 2



2 valence electrons

Group 17



7 valence electrons

Group 16, tellurium 54A and has 6 valence electrons. How many valence electrons will tellurium have?

How many electrons?

Hydrogen and Helium are both in Group 1A. Hydrogen 1 valence electron. Helium 2 valence electrons.

**Group 18A:** - These are noble gases that have a full shell of electrons. Elements are in Group 18 and 18A are noble gases. Helium, neon, argon, krypton, xenon, and radon are noble gases. "He 2, Ne 10, Ar 18, Kr 36, Xe 54, Rn 86" Group 18A has a full shell of valence electrons. Helium 2, Neon 10, Argon 18, Krypton 36, Xenon 54, Radon 86.

# Periods And Groups Valence Electrons Masses Answers

**Xiaolong Qi**



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