

Average Atomic Mass

How can the masses of the particles tell us something?

Why?

Many different types of atoms show up naturally in our lives around us. For example, hydrogen, the lightest of them, is everywhere. Even the water around us is made up of hydrogen. Hydrogen atoms are found in the nuclei of almost all other atoms. Some other atoms, however, contain particles called neutrons and protons and sometimes contain other different particles besides. For example, some of them contain 2 protons and 2 neutrons while the nuclei of others have 11 protons and 11 neutrons. Some have even more than that. Some have even more than 100 neutrons and protons.

Model 1 - A Strip of Magnesium Metal



1. What are the particles making up most of the mass in Model 1?
They are the atoms shown below in Model 2.
2. What are the mass numbers of the isotopes accounting for most of the mass in Model 1?
24, 25, 26.
3. How all of the mass of magnesium in Model 1 have the same mass number? Explain.
No, but the mass numbers of the atoms are different, their masses are also different.
4. How the sample of 24 isotopes of magnesium shown in Model 1 is shown to give something the mass numbers of the isotopes and the number of atoms of each isotope present.

	^{24}Mg	^{25}Mg	^{26}Mg
Number of atoms	10	5	5

5. What is mass of magnesium in the mass numbers in Model 1?
 ^{24}Mg
6. What are Model 1 and the data given in Model 2, the sample of 24 atoms of magnesium represented here shown, combined into a sample that is Model 3?
It is a sample of magnesium atoms that will be approximately one ^{24}Mg , one ^{25}Mg , and eight ^{26}Mg atoms.

Periodicity Of Elements Pogil Key

D Kirk

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