Chemistry I: Periodicity Worksheet

1) Arrange the following elements in order of increasing size: Al, B, C, K, and Na.

C B Al Na K smallest largest

Arrange the following elements in order of increasing size: Ca, Rb, P, Ge, and Sr.

P Ge Ca Sr Rb smallest largest

Circle the atom or ion in each pair that has the larger radius.

- a) Clor CI
- b) All or O
- c) In or I

4) Select the atom or ion in each pair that has the larger radius.

- a) Cs or Rb
- b) O<sup>3</sup> or O
- c) BrorAs

5) Which of the following groups of elements is arranged correctly in order of increasing ionization energy? (Answer by circling the correct answer.)

Non

- a)  $C \le Si \le Li \le Ne$
- b) Ne  $\leq$  Si  $\leq$  C  $\leq$  Li
- e) Li < Si < C < Ne
- d) Ne < C < Si < Li</p>

6) Arrange the following atoms in order of increasing ionization energy: Li, K, C, and N.

K LI C N lowest highest

Arrange the following atoms in order of increasing ionization energy: Si, K, As, and Ca.

K Ca As Si lowest highest

8) Compare the elements Na, Mg, O, and P.

- a) Which has the largest atomic radius?
- b) Which has the highest electron affinity? O

c) Place the elements in order of increasing ionization energy.

Na Mg P O lowest highest

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Ying Liu

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# **Table of Contents Periodicity Worksheet Answer Key**

- 1. Understanding the eBook Periodicity Worksheet Answer Key
  - The Rise of Digital Reading Periodicity Worksheet Answer Key
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Periodicity Worksheet Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodicity Worksheet Answer Key
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Periodicity Worksheet Answer Key
  - Personalized Recommendations
  - Periodicity Worksheet Answer Key User Reviews and Ratings

- Periodicity Worksheet Answer Key and Bestseller Lists
- 5. Accessing Periodicity Worksheet Answer Key Free and Paid eBooks
  - Periodicity Worksheet Answer Key Public Domain eBooks
  - Periodicity Worksheet Answer Key eBook Subscription Services
  - Periodicity Worksheet Answer Key Budget-Friendly Options
- 6. Navigating Periodicity Worksheet Answer Key eBook Formats
  - o ePub, PDF, MOBI, and More
  - Periodicity Worksheet Answer Key Compatibility with Devices
  - Periodicity Worksheet Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodicity Worksheet Answer Key
  - Highlighting and Note-Taking Periodicity Worksheet Answer Key
  - Interactive Elements Periodicity Worksheet Answer Key
- 8. Staying Engaged with Periodicity Worksheet Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodicity Worksheet Answer Key
- 9. Balancing eBooks and Physical Books Periodicity Worksheet Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodicity Worksheet Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Periodicity Worksheet Answer Key
  - Setting Reading Goals Periodicity Worksheet Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodicity Worksheet Answer Key
  - Fact-Checking eBook Content of Periodicity Worksheet Answer Key
  - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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