

PERIODIC CLASSIFICATION OF ELEMENT

DOBERNIER TRIADS:-

- Arranged elements in the group of 3.
- Atomic mass of middle element is average of atomic mass of other two elements.



Drawbacks:-

- as known elements were not grouped

Lothar Meyer's Curves:-



- Alkali metals → top position
- Alkaline earth metals → descending
- Halogen → ascending position.

MODERN PERIODIC TABLE:-

- long form of periodic table
- based on Moseley law: $\sqrt{f} = a - b(Z - c)$
- Periodic properties are periodic function of their atomic number.

Periodicity - Repetition of similar properties of elements after regular interval of time.

Cause of periodicity - due to similar e^- configuration.

Transition elements: they have incompletely filled d-orbitals in their ground state.

Zn, Cd, Hg → are not transition elements because of completely filled d-orbitals. (pseudo transition elements).

Representative elements - Combined elements of s and p block are known as representative elements.

NEWLAND'S OCTAVE RULE:-

- When elements are arranged in increasing order of atomic mass, then every 8th element standing from first will show same property.



Drawbacks:-

1. Valid upto calcium only
2. No position of inert gases
3. Not worked on heavier elements.

MENDLEEV'S PERIODIC

- Periodic property of elements are periodic function of their atomic masses.

Merits:-

- Simplified study
- Some vacant space was left for the elements which were not known at that time.

Eka aluminium → Ga
Eka silicon → Ge

- Corrected atomic mass of element

Demerits:-

- (a) Position of hydrogen was not fixed
- (b) Could not explain periodicity.

MODERN PERIODIC TABLE:-

18 groups and 7 periods.

4-Blocks

s-Block = G_1 and $G_2 = ns^1-2$

p-Block = $G_{13} - G_{18} = ns^2 np^1-6$

d-Block = $G_3 - G_{10} = (n-1)d^1-10 ns^2$

f-Block = lanthanoid = $2 = 58 - 71$

actinoid = $2 = 90 - 103$

G_3 and P_6

Diagonal Relation;



Nomenclature;

0 → nil	5 → pent	Suffix = turn
1 → un	6 → hex	
2 → bi	7 → sept	100 unilipentium
3 → tri	8 → oct	
4 → quad	9 → non	

EFFECTIVE NUCLEAR CHARGE

$$Z_{\text{eff}} = Z - \sigma$$

Trends - Period → $Z_{\text{eff}} \uparrow$

Group → Z_{eff} almost remains same. or $Z_{\text{eff}} \uparrow \downarrow$

Shielding → $s > p > d > f$

- more +ve charge, more is Z_{eff}
- more -ve charge, less is Z_{eff}

IONISATION ENERGY

- Energy required to remove an electron from an isolated gaseous atom.



$$\text{IE} = \text{IE}_1 < \text{IE}_2 < \text{IE}_3 \dots$$

Trends - $Z_{\text{eff}} \propto \text{IE}$

Period - $Z_{\text{eff}} \uparrow = \text{IE} \uparrow$

Group - $Z_{\text{eff}} \downarrow = \text{IE} \downarrow$

$G_2 = \text{O} > \text{F} > \text{N} > \text{C} > \text{B} > \text{Be} > \text{Li}$

$G_{17} = \text{Cs} > \text{Rb} > \text{K} > \text{Na} > \text{Li}$

Period 2:-



ATOMIC RADIUS:-

- Distance b/w nucleus and electron of outermost shell.
- But it's impossible to find exact atomic radii.

Reason-1 = atoms too small

2 = atom doesn't have any boundary

Covalent Radius - Half of the internuclear distance b/w 2 atoms bonded by covalent bond.



Metallic Radius - Half of the internuclear distance b/w 2 atoms bonded by a metallic bond.



Van der Waals Radius - Half of the internuclear distance b/w 2 nearest molecules in non-bonded state.



Trends: Atomic Radius $\propto \frac{1}{Z_{\text{eff}}}$
Period → $Z_{\text{eff}} \uparrow = \text{Atomic Radius} \downarrow$
Group → $Z_{\text{eff}} \downarrow = \text{Atomic Radius} \uparrow$

Ionic Radius + Radius < Neutral Radius

ELECTRON GAIN ENTHALPY:-

- Energy required or released when an e^- is added to outermost orbit of an isolated gaseous atom. It is known as Electron gain enthalpy.



Note - In majority cases $\Delta_{\text{eg}} H_1$ is -ve (exothermic)

Exception = N } $\Delta_{\text{eg}} H_1 = +ve$ due to stable configuration

Note - Successive electron gain enthalpy is always +ve.

Trends: $\Delta_{\text{eg}} H \propto Z_{\text{eff}}$
Period - $Z_{\text{eff}} \uparrow \Rightarrow \Delta_{\text{eg}} H \uparrow$
Group - $Z_{\text{eff}} \downarrow \Rightarrow \Delta_{\text{eg}} H \downarrow$

$G_2 = \text{S} > \text{Se} > \text{Te} > \text{Po} > \text{O}$
→ due to small size Reason!
→ e^-e^- repulsion.

Periodic Classification 3rd Chapter 11th Class

L Reisser



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