

# Graphing Periodic Trends Activity

Name: \_\_\_\_\_ Period: \_\_\_\_\_

Use the information in these tables to complete the graph as described below

| Symbol | Atomic Radius (Picometers) | First Ionization energy (kilojoules/mole) | Electronegativity (4-point scale) |
|--------|----------------------------|-------------------------------------------|-----------------------------------|
| H      | 31                         | 1312                                      | 2.1                               |
| Li     | 128                        | 520                                       | 1.0                               |
| Na     | 166                        | 496                                       | 0.9                               |
| K      | 203                        | 410                                       | 0.8                               |
| Rb     | 220                        | 403                                       | 0.8                               |
| Cs     | 244                        | 376                                       | 0.7                               |

| Symbol | Atomic Radius (Picometers) | First Ionization energy (kilojoules/mole) | Electronegativity (4-point scale) |
|--------|----------------------------|-------------------------------------------|-----------------------------------|
| Na     | 166                        | 496                                       | 0.9                               |
| Mg     | 141                        | 738                                       | 1.2                               |
| Al     | 121                        | 578                                       | 1.5                               |
| Si     | 111                        | 787                                       | 1.8                               |
| P      | 107                        | 1012                                      | 2.1                               |
| S      | 105                        | 1000                                      | 2.5                               |
| Cl     | 102                        | 1251                                      | 3.0                               |
| Ar     | 106                        | 1521                                      | ---                               |

- Using colored pencils or pen, list under the symbol (in this order!) the Atomic Radius, First Ionization Energy, and Electronegativity.
- Use a different color for each property. Example: Write all of the Atomic radius values in red, all of the First ionization energies in green, and all of the Electronegativities in blue. Of course, you can pick any colors that are available, as long as you are consistent.
- Observe the trends in each property as you go down the Alkali metal group, and as you go across Period 3.
- Write out each of the statements written on the opposite side, completing each statement with the observed trend (increase or decrease)

|                         |                                                                                                                                                                                          |    |    |   |   |    |    |
|-------------------------|------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|----|----|---|---|----|----|
| H                       | <p>Across a period (→) atomic radius tends to _____.</p> <p>Across a period (→) first ionization energy tends to _____.</p> <p>Across a period (→) electronegativity tends to _____.</p> |    |    |   |   |    |    |
| Li                      |                                                                                                        |    |    |   |   |    |    |
| Na<br>166<br>496<br>0.9 | Mg                                                                                                                                                                                       | Al | Si | P | S | Cl | Ar |
| K                       | <p>Down a Group (↓) atomic radius tends to _____.</p> <p>Down a Group (↓) first ionization energy tends to _____.</p> <p>Down a Group (↓) electronegativity tends to _____.</p>          |    |    |   |   |    |    |
| Rb                      |                                                                                                                                                                                          |    |    |   |   |    |    |
| Cs                      |                                                                                                                                                                                          |    |    |   |   |    |    |

# Periodic Trends Answers Key

**Rachel Sandford**



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