

\* %yield could be in **g** or **mol**

$$\% \text{yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

$$= \frac{7.54\text{g CaO}}{8.97\text{g CaO}} \times 100 = 0.841 \times 100$$
$$= \boxed{84.1 \%}$$

# Percent Yield Holt Chemfile Answers

**Ralph Thomas Myers**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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