

Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

Remember: Limiting reagent is the reactant that will completely run out first in a chemical reaction.
Calculate: Molar masses, stoichiometric coefficients, and percent yield. Remember: Limiting reagent is the reactant that will completely run out first in a chemical reaction.

Remember: Limiting reagent is the reactant that will completely run out first in a chemical reaction.
 Calculate: Molar masses, stoichiometric coefficients, and percent yield. Remember: Limiting reagent is the reactant that will completely run out first in a chemical reaction.



The reaction is balanced. The coefficients are 2, 1, and 2.

2H ₂	2 x 2.016 g/mol
O ₂	2 x 16.00 g/mol
2H ₂ O	2 x 18.016 g/mol
2H ₂	2 x 2.016 g/mol

1. 2.016 g/mol is the molar mass of H₂. 2.016 g/mol is the molar mass of O₂. 18.016 g/mol is the molar mass of H₂O.



2. 2.016 g/mol is the molar mass of H₂. 2.016 g/mol is the molar mass of O₂. 18.016 g/mol is the molar mass of H₂O.

3. 18.016 g/mol is the molar mass of H₂O. 18.016 g/mol is the molar mass of H₂O. 18.016 g/mol is the molar mass of H₂O.



4. 18.016 g/mol is the molar mass of H₂O. 18.016 g/mol is the molar mass of H₂O. 18.016 g/mol is the molar mass of H₂O.



Percent Yield And Limiting Reactants Quiz

Paul E. Curtis, Roger L. Courson



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web i greet you in the name of our lord jesus let me take this opportunity the lord has granted to us to come and worship him during this day that we have never seen in our lives we want to take this chance to glorify his name for the grace that has enabled us to meet here for this program of the day

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