

13.1

THE NATURE OF GASES

Section Review

Objectives

1. describe the characteristics of ideal gases and deviations
2. distinguish between an ideal and real gas
3. define the macroscopic terms appropriate to gaseous states

Vocabulary

- | | | |
|---------------|-------------------------|------------------|
| a. ideal gas | b. pressure | c. pressure |
| d. atmosphere | e. atmospheric pressure | f. absolute zero |
| g. gas laws | h. volume | |

Part A Completion

Use the vocabulary terms in parentheses to complete each of the sentences. Write the complete sentence on the lines below.

Pressure is the force exerted by _____, depending on the number of _____ molecules striking the _____ the force exerted by a single molecule has a value _____ the force exerted by a group of molecules. _____, _____, and _____ are all states of matter. _____ matter has _____ characteristics. _____ states matter what _____ pressure is the _____ exerted by _____ particles contained in a _____ container and is _____ by the _____ and _____ of the _____ and is _____ by _____.

1	_____
2	_____
3	_____
4	_____
5	_____
6	_____
7	_____
8	_____
9	_____
10	_____

Part B True-False

Mark each of the following statements as true (T) or false (F). Write true or false.

- _____ 11. Gases are composed of particles moving rapidly.
- _____ 12. Gases are composed of particles moving.

Section 1 The Nature Of Gases Answers

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