Chapter 4 Atomic Structure

Section 4.2 The Structure of an Atom (pages 108–112)

This section compares the properties of three subatomic particles. It also discusses atomic numbers, mass numbers, and isotopes.

Reading Strategy (page 108)

Monitoring Your Understanding Before you read, list in the table shown what you know about atoms and what you would like to learn. After you read, list what you have learned. For more information on this Reading Strategy, see the Reading and Study Skills in the Skills and Reference Handbook at the end of your textbook.

What I Know	What I Would	What I Have
About Atoms	Like to Learn	Learned
Most students will know that atoms are the "building blocks" of matter, and some may know that atoms contain subatomic particles.	Based on the title of the section, students may say that they want to learn more about the structure of atoms.	

Properties of Subatomic Particles (pages 108-109)

1. What are three subatomic particles?

a. Protons b. Electrons c. Neutrons

Circle the letter that identifies a subatomic particle with a positive charge.

a. nucleus

b) proton

c. neutron

d. electron

Why did Chadwick conclude that the particles produced by his experiment were neutral in charge? A charged object did not deflect the paths of the particles.

Comparing Subatomic Particles (pages 109-110)

Circle the letters of properties that vary among subatomic particles.

a. color

(b) mass

c. charge

- d location in the atom
- Circle the letter of the expression that accurately compares the masses of neutrons and protons.
 - a mass of 1 neutron = mass of 1 proton
 - b. mass of 2000 neutrons = mass of 1 proton
 - c. mass of 1 electron = mass of 1 proton
 - d. mass of 1 neutron = mass of 1 electron

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Table of Contents Section 1 Studying Atoms Answers

- 1. Understanding the eBook Section 1 Studying Atoms Answers
 - The Rise of Digital Reading Section 1 Studying Atoms Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Section 1 Studying Atoms Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 1 Studying Atoms Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Section 1 Studying Atoms Answers

- Personalized Recommendations
- Section 1 Studying Atoms Answers User Reviews and Ratings
- Section 1 Studying Atoms Answers and Bestseller Lists
- 5. Accessing Section 1 Studying Atoms Answers Free and Paid eBooks
 - Section 1 Studying Atoms Answers Public Domain eBooks
 - Section 1 Studying Atoms Answers eBook Subscription Services
 - Section 1 Studying Atoms Answers Budget-Friendly Options
- 6. Navigating Section 1 Studying Atoms Answers eBook Formats
 - o ePub, PDF, MOBI, and More
 - Section 1 Studying Atoms Answers Compatibility with Devices
 - Section 1 Studying Atoms Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 1 Studying Atoms Answers
 - Highlighting and Note-Taking Section 1 Studying Atoms Answers
 - Interactive Elements Section 1 Studying Atoms Answers
- 8. Staying Engaged with Section 1 Studying Atoms Answers
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 1 Studying Atoms Answers
- 9. Balancing eBooks and Physical Books Section 1 Studying Atoms Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 1 Studying Atoms Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Section 1 Studying Atoms Answers
 - Setting Reading Goals Section 1 Studying Atoms Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section 1 Studying Atoms Answers

- Fact-Checking eBook Content of Section 1 Studying Atoms Answers
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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