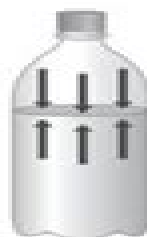


15. Describe the two diagrams of a bottled carbonated beverage below as *greater pressure* or *lower pressure*, and then as *greater solubility* or *lower solubility*. How do these two examples illustrate the relationship between the solubility of a gas and its vapor pressure?

In the closed bottle, the pressure is high and lots of CO_2 is dissolved in the
carbonated beverage. Once the cap is removed, the vapor pressure decreases,
allowing more CO_2 in the carbonated beverage to escape.



greater pressure and greater solubility lower pressure and lower solubility

16. How does a solution become supersaturated? If the temperature of a solution having
a small excess of solid solute is raised, the solute dissolves. If the solution is then
allowed to cool slowly, the excess solute may stay dissolved at a temperature below
the temperature at which it would ordinarily crystallize.

SECTION 16.2 CONCENTRATIONS OF SOLUTIONS (pages 480–486)

This section explains how to solve problems involving molarity of a solution, how to prepare dilute solutions from more concentrated solutions, and what is meant by percent by volume and percent by mass.

► Molarity (pages 480–482)

1. A measure of the amount of solute dissolved in a given quantity of solvent is the concentration of a solution.
2. The most important unit of concentration in chemistry is molarity.
3. Is the following sentence true or false? Molarity is the number of moles of dissolved solute per liter of solvent. false

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Silberberg



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General Chemistry Donald A. McQuarrie, Stanley Gill, 2011-06-15 This Fourth Edition of McQuarrie's classic text offers a thorough revision and a quantum leap forward from the previous edition Taking an atoms first approach it promises to be another ground breaking text in the tradition of McQuarrie's many previous works This outstanding new text available in a soft cover edition offers professors a fresh choice and outstanding value

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The Pearson Guide to Physical Chemistry for the IIT JEE Singhal, **Science Activity Book Chapterwise Class 12** Priti Singhal, 2024-11-17. This book is designed to ignite curiosity and foster a love for science in students from grades 1 to 12. With a diverse range of engaging activities, this book aims to provide a hands-on, interactive approach to understanding fundamental scientific concepts tailored to the unique developmental stages across all grade levels. Our primary goal is to make learning science enjoyable and enriching. The book is filled with colourful illustrations, real life examples, and interactive exercises that help students understand and relate to the world around them. Each chapter is carefully structured to build on prior knowledge, ensuring a steady progression in learning as students advance through the grades.

Fundamentals of Physical Chemistry Ananya Ganguly, 2011. Fundamentals of Physical Chemistry is the signature compilation of the class-tested notes of iconic chemistry coach Ananya Ganguly. Her unique teaching methodology and authoritative approach in teaching of concepts, their application, and strategy is ideal for preparing for the IITJEE examinations. The author's impeccable command and the authority on each foray of chemistry teaching are visible in each chapter, and the chapter-ending exercises. Each chapter unfolds the structured, systematic, and patterned chemistry concepts in a lucid and student-friendly approach. The book is without those unnecessary frills that make the bulk in other popular books in the market for the IITJEE. An indispensable must-have for in-depth comprehension of Chemistry for the coveted IITJEE.

Physical Chemistry for Engineering and Applied Sciences Frank R. Foulkes, 2012-09-12. Physical Chemistry for Engineering and Applied Sciences is the product of over 30 years of teaching first-year Physical Chemistry as part of the Faculty of Applied Science and Engineering at the University of Toronto. Designed to be as rigorous as compatible with a first-year student's ability to understand the text, it presents detailed step-by-step derivations of the equations that permit the student to follow the underlying logic and of equal importance to appreciate any simplifying assumptions made or mathematical tricks employed. In addition to the 600 exercises and end-of-chapter problems, the text is rich in worked non-trivial examples, many of which are designed to be inspiring and thought-provoking. Step-by-step derivation of all equations

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we represent a generic sequence as $a_1, a_2, a_3, \dots, a_n$ in order to define a sequence we must give enough information to find its n th term two ways of doing this are 1 with a formula e.g. $a_n = 1$

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