

Section 1 Describing Chemical Reactions

A **chemical reaction** is the process by which one or more substances are changed into one or more different substances.

In any chemical reaction, the **original substances** are known as the **reactants**, and the **resulting substances** are known as the **products**.

According to the **law of conservation of mass**, the **total mass of reactants** must equal the **total mass of products** for any given chemical reaction.

Chemical reactions are described by **chemical equations**.

A **chemical equation** represents, with symbols and formulas, the **identities and relative molecular or molar amounts** of the **reactants and products** in a chemical reaction.

Indications of a Chemical Reaction

- **Evolution of energy as heat and light** is strong evidence that a chemical reaction has taken place.
Example, the decomposition of ammonium dichromate is accompanied by the evolution of energy as heat and light.
chemical reaction occurs between natural gas and oxygen if you burn gas for cooking in your house.
- **Production of a gas** the evolution of gas bubbles when two substances are mixed is often evidence of a chemical reaction.
Example, bubbles of carbon dioxide gas form immediately when baking soda is mixed with vinegar.
- **Formation of a precipitate**
a **precipitate** is a solid that is produced as a result of a chemical reaction in solution and that separates from the solution.
- **Color change** is often an indication of a chemical reaction.

Characteristics of Chemical Equations

The following requirements will aid you in writing and reading chemical equations correctly.

1. The equation must represent known facts.
 2. The equation must contain the correct formulas for the reactants and products.
 3. The law of conservation of mass must be satisfied.
- A **coefficient** is a small whole number that appears in front of a formula in a chemical equation.

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