

Table 1. Types of Jacobsthal quaternions [20–23].

	Definition	Multiplication Rules
Jacobsthal quaternions	$JQ_n = (J_n, J_{n+1}, J_{n+2}, J_{n+3})$ $J_n = J_{n-1} + 2J_{n-2}, J_1 = J_2 = 1$	$(1, i, j, k), i^2 = j^2 = k^2 = -1$ $ij = -ji = k, jk = -kj = i, ki = -ik = j$
k-jacobsthal quaternions	$QJ_{k,n} = (J_{k,n}, J_{k,n+1}, J_{k,n+2}, J_{k,n+3})$ $QJ_{k,n+2} = kQJ_{k,n+1} + 2QJ_{k,n}$	$(1, i_1, i_2, i_3), i_1^2 = i_2^2 = i_3^2 = i_1 i_2 i_3 = -1$
Dual Jacobsthal quaternions	$D_n^J = (J_n, J_{n+1}, J_{n+2}, J_{n+3})$ $J_n = J_{n-1} + 2J_{n-2}, J_1 = J_2 = 1$	$(1, i, j, k) i^2 = j^2 = k^2 = ijk = 0$ $ij = -ji = jk = -kj = ki = -ik = 0$
Dual-complex Jacobsthal quaternions	$QJ_n = (J_n, J_{n+1}, J_{n+2}, J_{n+3})$ $J_n = J_{n-1} + 2J_{n-2}, J_1 = J_2 = 1$	$(1, i, \epsilon, i\epsilon), i^2 = -1, \epsilon^2 = 0, \epsilon \neq 0$ $(i\epsilon)^2 = 0$

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Puja Mehta



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web 1. molarity $M = \frac{\text{mass}}{\text{molar mass} \times \text{volume}}$ $\frac{0.4500 \text{ g}}{0.0825 \text{ L} \times 141.9579 \text{ g/mol}} = 0.0129 \text{ M}$ 2. molality $m = \frac{\text{mass}}{\text{molar mass} \times \text{mass of solvent}}$ $\frac{0.825 \text{ g}}{0.00581158 \text{ kg} \times 141.9579 \text{ g/mol}} = 0.00581158 \text{ m}$ 3. mole fraction $n_2 = \frac{0.4500 \text{ g}}{0.0129 \text{ M} \times 141.9579 \text{ g/mol}} = 0.00581158 \text{ mol}$ 4. $\frac{0.825 \text{ g}}{0.00581158 \text{ mol} \times 141.9579 \text{ g/mol}} = 0.00581158 \text{ mol}$ 5. $\frac{0.4500 \text{ g}}{0.0129 \text{ M} \times 141.9579 \text{ g/mol}} = 0.00581158 \text{ mol}$ 6. $\frac{0.825 \text{ g}}{0.00581158 \text{ mol} \times 141.9579 \text{ g/mol}} = 0.00581158 \text{ mol}$ 7. $\frac{0.4500 \text{ g}}{0.0129 \text{ M} \times 141.9579 \text{ g/mol}} = 0.00581158 \text{ mol}$ 8. $\frac{0.825 \text{ g}}{0.00581158 \text{ mol} \times 141.9579 \text{ g/mol}} = 0.00581158 \text{ mol}$ 9. $\frac{0.4500 \text{ g}}{0.0129 \text{ M} \times 141.9579 \text{ g/mol}} = 0.00581158 \text{ mol}$ 10. $\frac{0.825 \text{ g}}{0.00581158 \text{ mol} \times 141.9579 \text{ g/mol}} = 0.00581158 \text{ mol}$ 11. $\frac{0.4500 \text{ g}}{0.0129 \text{ M} \times 141.9579 \text{ g/mol}} = 0.00581158 \text{ mol}$ 12. $\frac{0.825 \text{ g}}{0.00581158 \text{ mol} \times 141.9579 \text{ g/mol}} = 0.00581158 \text{ mol}$ 13. $\frac{0.4500 \text{ g}}{0.0129 \text{ M} \times 141.9579 \text{ g/mol}} = 0.00581158 \text{ mol}$ 14. $\frac{0.825 \text{ g}}{0.00581158 \text{ mol} \times 141.9579 \text{ g/mol}} = 0.00581158 \text{ mol}$ 15. $\frac{0.4500 \text{ g}}{0.0129 \text{ M} \times 141.9579 \text{ g/mol}} = 0.00581158 \text{ mol}$

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web apr 7 2014 molarity moles of solute / liters of solution 8.422 a first convert 250 ml to liters 250 / 1000 = 0.25 then calculate molarity 5 moles / 0.25 liters = 20 M 3. c a solution with molarity 2 requires 2 M of NaOH per liter so 4 x 2.8 M = 11.2 M a solution of molarity 1.5 M requires 1.5 mol of NaOH to every litre of solvent

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web answer the correct option is c explanation the molarity of a mixture can be calculated using the following formula $M_{\text{mix}} = \frac{M_1 V_1 + M_2 V_2 + \dots + M_n V_n}{V_1 + V_2 + \dots + V_n}$ $\frac{1.5 \times 480 + 2 \times 520 + 3 \times 480 + 4 \times 520}{480 + 520 + 480 + 520} = 2.5$ 1000 / 344 = 2.91 M 5. with increase in temperature which of these changes a molality b fraction of solute present in water

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web 1. calculate the molarity of a solution which contains 0.40 mol of a substance dissolved in 1.6 L of a solution 2. what is the molarity of a solution containing 325 g of NaCl dissolved in 750 ml of solution 3. 140 g of KCl is dissolved in 600 ml of water what is the molarity 4. 724.4 g of ammonium phosphate in 4500 ml of alcohol

6.1 calculating molarity problems chemistry libretexts - Jun 13 2023

web determine the molarity for each of the following solutions 0.444 mol of COCl_2 in 0.654 L of solution 98.0 g of phosphoric acid H_3PO_4 in 1.00 L of solution 0.2074 g of calcium hydroxide Ca(OH)_2 in 40.00 ml of solution 10.5 kg

16.11 molality chemistry libretexts - Aug 15 2023

web molality differs from molarity only in the denominator while molarity is based on the liters of solution molality is based on the kilograms of solvent concentrations expressed in molality are used when studying properties of solutions related to vapor pressure and temperature changes

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web molarity molality notes and practice answer the questions below show all work including units watch your significant digits and circle your answers molarity just a reminder molarity is one of the many ways to measure concentration or