

Specific Heat

DIRECTIONS: Use formula $q = (m)(\Delta T)(C)$ to solve the following problems. Show all work and units. Use <http://weebly.com> for specific heat of various substances.

1. A 30g piece of metal absorbs 1,200 joules of heat energy and its temperature change from 25°C to 175°C. Calculate the specific capacity of the metal. What is the likely metal?
2. How many joules of heat are needed to raise the temperature of 15.0 g of iron from 11°C to 68°C, if the specific heat of iron is 0.45 J/g°C?
3. To what temperature will a 50.0 g piece of brass raise if it absorbs 6 kilojoules of heat and its specific heat capacity is 0.38 J/g°C? The initial temperature of the brass is 20.0°C.
4. Identify the substance if 1500.0 g of the substance absorbs 74,500 joules of heat, and its temperature changes from 25°C to 47°C.
5. 100.0 g of 1.0°C water is heated until its temperature is 37°C. If the specific heat of water is 4.18 J/g°C, calculate the amount of heat energy needed to cause this rise in temperature.
6. 25.0 g of antimony is heated from 30°C to 160°C, and absorbs 750 joules of heat in the process. Calculate the specific heat capacity of antimony.
7. What is the specific heat capacity of quartz glass if 200.0 g of the glass absorbs 2520 joules of heat and the temperature rises 18.0°C?
8. If a sample of titanium is initially at 25°C, what is its final temperature if 250.0 g of titanium absorbs 1.0 kilojoules of heat, and the specific heat of titanium is 0.523 J/g°C?
9. How much energy would be absorbed by 300.0 g of steel to increase its temperature from 100.0°C to 450.0°C? (C of steel = .48 J/g °C)
10. When 55.0 g of platinum drops its temperature from 300.0°C to 250.0°C, how much heat energy is released? (C of platinum = .133 J/g °C)
11. How much energy is required to heat 125.0 g of lithium from 2.0°C to 30.0°C? (C of lithium = 3.6 J/g°C)
12. How much heat (in kJ) is given out when 350.0 g of lead cools from 200.0°C to 30.0°C?

Practice Problems For Specific Heat

ML Yell



Practice Problems For Specific Heat:

Adopting the Beat of Expression: An Mental Symphony within **Practice Problems For Specific Heat**

In some sort of taken by displays and the ceaseless chatter of instant communication, the melodic beauty and psychological symphony developed by the published term often disappear in to the backdrop, eclipsed by the constant noise and disruptions that permeate our lives. But, nestled within the pages of **Practice Problems For Specific Heat** a stunning literary value full of raw feelings, lies an immersive symphony waiting to be embraced. Constructed by a wonderful musician of language, that charming masterpiece conducts readers on an emotional trip, well unraveling the hidden songs and profound affect resonating within each cautiously constructed phrase. Within the depths of this moving review, we shall explore the book is central harmonies, analyze its enthralling publishing style, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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