

9. Each of four test tubes contains a different concentration of HCl(aq) at 25°C . A 1-gram cube of Zn is added to each test tube. In which test tube is the reaction occurring at the slowest rate?

A) 1.0 M HCl(aq)



B) 0.1 M HCl(aq)



C) 0.01 M HCl(aq)



D) 0.001 M HCl(aq)



10. As the concentration of reacting particles increases, the rate of reaction generally
- A) decreases B) increases
C) remains the same
11. Which change would most likely decrease the rate of a chemical reaction?
- A) decreasing a reactant's concentration
B) decreasing a reactant's surface area
C) cooling the reaction mixture
D) adding a catalyst to the reaction mixture
12. Which statement best describes how a catalyst increases the rate of a reaction?
- A) The catalyst provides an alternate reaction pathway with a higher activation energy.
B) The catalyst provides an alternate reaction pathway with a lower activation energy.
C) The catalyst provides the same reaction pathway with a higher activation energy.
D) The catalyst provides the same reaction pathway with a lower activation energy.
13. If the pressure on gaseous reactants is increased, the rate of reaction is increased because there is an increase in the
- A) activation energy B) volume
C) concentration D) heat of reaction

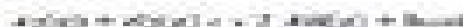
14. A 1.0-gram piece of Zn reacts with 1 milliliter of HCl(aq) . Which of these conditions of concentration and temperature would produce the greatest rate of reaction?

- A) 1.0 M HCl(aq) at 20°C
B) 1.0 M HCl(aq) at 40°C
C) 2.0 M HCl(aq) at 20°C
D) 2.0 M HCl(aq) at 40°C

15. In a biochemical reaction, an enzyme acts as a catalyst. Which the

- A) activation energy of the reaction to decrease
B) potential energy of the products to decrease
C) kinetic energy of the reactants to increase
D) heat of reaction to increase

16. Which reaction



As increase in the concentration of $\text{A}_2(\text{g})$ will

- A) decrease the production of $\text{AB}(\text{g})$
B) decrease the frequency of collisions between $\text{A}_2(\text{g})$ and $\text{B}_2(\text{g})$
C) increase the production of $\text{AB}(\text{g})$
D) increase the frequency of collisions between $\text{A}_2(\text{g})$ and $\text{B}_2(\text{g})$

17. A reaction is most likely to occur when reactant particles collide with

- A) proper energy, only
B) proper orientation, only
C) both proper energy and proper orientation
D) neither proper energy nor proper orientation

18. Given the reaction



The reaction will most likely occur at the greatest rate if A and B represent

- A) ionic molecular compounds in the solid phase
B) ionic compounds in the solid phase
C) volatile or nonpolar molecular compounds
D) volatile or ionic compounds

Practice Packet Unit 8 Kinetics And Equilibrium Answers

Victor M. Corman



Practice Packet Unit 8 Kinetics And Equilibrium Answers:

This book delves into Practice Packet Unit 8 Kinetics And Equilibrium Answers. Practice Packet Unit 8 Kinetics And Equilibrium Answers is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Practice Packet Unit 8 Kinetics And Equilibrium Answers, encompassing both the fundamentals and more intricate discussions.

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