

Practice Problems

Do your best to solve all the following problems.

SECTION 3.1 THE MOLE: A MEASUREMENT OF MATTER

1. Calculate the molar mass of each of the following compounds:

a. C_2H_6 b. $\text{C}_2\text{H}_4\text{O}$ c. $\text{C}_2\text{H}_5\text{OH}$ d. $\text{C}_2\text{H}_5\text{NH}_2$ e. $\text{C}_2\text{H}_5\text{COOH}$

2. Calculate the mass of each of the following compounds:

a. 0.500 mol C_2H_6 b. 0.250 mol $\text{C}_2\text{H}_4\text{O}$

c. 0.125 mol $\text{C}_2\text{H}_5\text{OH}$ d. 0.0625 mol $\text{C}_2\text{H}_5\text{NH}_2$

e. 0.0312 mol $\text{C}_2\text{H}_5\text{COOH}$ f. 0.0156 mol C_2H_6 g. 0.00781 mol $\text{C}_2\text{H}_4\text{O}$

3. Calculate the mass of:

a. 0.500 mol C_2H_6 b. 0.250 mol $\text{C}_2\text{H}_4\text{O}$

c. 0.125 mol $\text{C}_2\text{H}_5\text{OH}$ d. 0.0625 mol $\text{C}_2\text{H}_5\text{NH}_2$ e. 0.0312 mol $\text{C}_2\text{H}_5\text{COOH}$

f. 0.0156 mol C_2H_6 g. 0.00781 mol $\text{C}_2\text{H}_4\text{O}$ h. 0.00391 mol $\text{C}_2\text{H}_5\text{OH}$

i. 0.00195 mol $\text{C}_2\text{H}_5\text{NH}_2$ j. 0.000976 mol $\text{C}_2\text{H}_5\text{COOH}$

SECTION 3.2 MOLE—MASS AND MOLE—VOLUME RELATIONSHIPS

1. Calculate the molar mass of each of the following compounds:

a. C_2H_6 b. $\text{C}_2\text{H}_4\text{O}$ c. $\text{C}_2\text{H}_5\text{OH}$ d. $\text{C}_2\text{H}_5\text{NH}_2$ e. $\text{C}_2\text{H}_5\text{COOH}$

2. Calculate the mass of each of the following:

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e. 0.0312 mol $\text{C}_2\text{H}_5\text{COOH}$ f. 0.0156 mol C_2H_6 g. 0.00781 mol $\text{C}_2\text{H}_4\text{O}$

h. 0.00391 mol $\text{C}_2\text{H}_5\text{OH}$ i. 0.00195 mol $\text{C}_2\text{H}_5\text{NH}_2$

j. 0.000976 mol $\text{C}_2\text{H}_5\text{COOH}$ k. 0.000488 mol $\text{C}_2\text{H}_5\text{COOH}$ l. 0.000244 mol $\text{C}_2\text{H}_5\text{COOH}$

3. Calculate the mass of each of the following compounds:

a. 0.500 mol C_2H_6 b. 0.250 mol $\text{C}_2\text{H}_4\text{O}$ c. 0.125 mol $\text{C}_2\text{H}_5\text{OH}$ d. 0.0625 mol $\text{C}_2\text{H}_5\text{NH}_2$

e. 0.0312 mol $\text{C}_2\text{H}_5\text{COOH}$ f. 0.0156 mol C_2H_6 g. 0.00781 mol $\text{C}_2\text{H}_4\text{O}$

h. 0.00391 mol $\text{C}_2\text{H}_5\text{OH}$ i. 0.00195 mol $\text{C}_2\text{H}_5\text{NH}_2$ j. 0.000976 mol $\text{C}_2\text{H}_5\text{COOH}$

k. 0.000488 mol $\text{C}_2\text{H}_5\text{COOH}$ l. 0.000244 mol $\text{C}_2\text{H}_5\text{COOH}$ m. 0.000122 mol $\text{C}_2\text{H}_5\text{COOH}$

n. 0.000061 mol $\text{C}_2\text{H}_5\text{COOH}$ o. 0.0000305 mol $\text{C}_2\text{H}_5\text{COOH}$ p. 0.0000153 mol $\text{C}_2\text{H}_5\text{COOH}$

q. 0.00000765 mol $\text{C}_2\text{H}_5\text{COOH}$ r. 0.00000383 mol $\text{C}_2\text{H}_5\text{COOH}$

s. 0.00000191 mol $\text{C}_2\text{H}_5\text{COOH}$ t. 0.000000955 mol $\text{C}_2\text{H}_5\text{COOH}$

4. Calculate the mass of each of the following compounds:

a. 0.500 mol C_2H_6 b. 0.250 mol $\text{C}_2\text{H}_4\text{O}$ c. 0.125 mol $\text{C}_2\text{H}_5\text{OH}$ d. 0.0625 mol $\text{C}_2\text{H}_5\text{NH}_2$

e. 0.0312 mol $\text{C}_2\text{H}_5\text{COOH}$ f. 0.0156 mol C_2H_6 g. 0.00781 mol $\text{C}_2\text{H}_4\text{O}$ h. 0.00391 mol $\text{C}_2\text{H}_5\text{OH}$

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o. 0.0000305 mol $\text{C}_2\text{H}_5\text{COOH}$ p. 0.0000153 mol $\text{C}_2\text{H}_5\text{COOH}$ q. 0.00000765 mol $\text{C}_2\text{H}_5\text{COOH}$

r. 0.00000383 mol $\text{C}_2\text{H}_5\text{COOH}$ s. 0.00000191 mol $\text{C}_2\text{H}_5\text{COOH}$ t. 0.000000955 mol $\text{C}_2\text{H}_5\text{COOH}$

Section 10 Chemical Quantities Practice Problems

Answers

ML Morrison

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

Section 10 Chemical Quantities Practice Problems Answers:

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