



Lab Access Card

Classifying Chemical Reactions

Directions: Write the type of chemical reaction in the space provided next to each equation. Name the product or the products of the reaction at the end.

Column I

1. _____ 1. A precipitate forms in a gas liquid collection over water (hydrogen, acetylene, propylene)
2. _____ 2. Two reactants exchange products forming another substance
3. _____ 3. A hydrocarbon explodes, creates an ionic compound
4. _____ 4. Chloroalkanes break down into hydrogen and alkenes
5. _____ 5. A type of redox reaction that produces heat and light

Directions: Write the correct type of chemical reaction in the space provided.

hydrogenation _____ 6. $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$

combustion _____ 7. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$

oxidation

synthesis reaction

_____ 8. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$

single displacement reaction _____ 9. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$

synthesis reaction

_____ 10. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$

_____ 11. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$

combustion

single displacement reaction

_____ 12. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$

redox reaction

the decomposition reaction

_____ 13. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$

combustion

the decomposition reaction

_____ 14. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$

single displacement reaction _____ 15. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$

the decomposition reaction

Section 2 Reinforcement Classifying Chemical Reactions

Camilla Rothe



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