

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization | 5. K_a | 6. K_b |
| 7. ionization constant | 8. pH | |

Key Equation

- | | |
|-----------------------|-----------------------|
| 1. $[H^+] = 10^{-pH}$ | 2. $pH = -\log [H^+]$ |
| 3. $pH = -\log [H^+]$ | 4. $[H^+] = 10^{-pH}$ |

Part A Completion

Write the equation and/or the numerical value(s) of the constant(s) that are indicated in the spaces provided. Be sure to include correct units when appropriate.

1. Water autoionizes to H^+ and hydroxide ions. The concentration of H^+ in the autoionization of water at 25°C is 1.0×10^{-7} mol/L.
2. The autoionization constant for water is 1.0×10^{-14} at 25°C.
3. The K_a of acetic acid is 1.8×10^{-5} .
4. The K_b of ammonia is 1.8×10^{-5} .
5. The pH of a 0.1 M solution of acetic acid is 2.87.
6. The pH of a 0.1 M solution of ammonia is 11.13.
7. The K_a of acetic acid is 1.8×10^{-5} .
8. The K_b of ammonia is 1.8×10^{-5} .
9. The pH of a 0.1 M solution of acetic acid is 2.87.
10. The pH of a 0.1 M solution of ammonia is 11.13.

1. K_a
2. K_b
3. K_a
4. K_b
5. K_a
6. K_b
7. K_a
8. K_b
9. K_a
10. K_b

Section 19 2 Hydrogen Ions And Acidity Answer Key

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Velavancorresponding



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