

Solutions Colligative Properties DCI

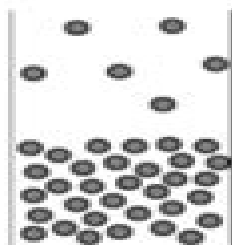
Name _____ Section _____

1. Define the term *colligative property* and list those physical properties of a solution that can be classified as colligative properties.

Colligative properties are properties which depend on the number of molecules or ions of solute present, and not on what the particles are (as long as they are not volatile). Properties of solutions which are colligative properties,

- 1) Vapor Pressure lowering
- 2) Boiling Point elevation
- 3) Freezing Point depression
- 4) Osmotic Pressure

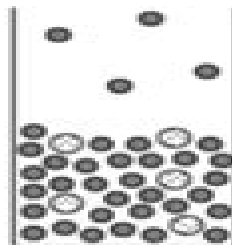
15. Illustrate and explain how the presence of a nonvolatile solute affects the vapor pressure of a liquid. See Appendix III for recommended demonstration, video, or computer resources.



● H_2O

Before addition of the solute

The vapor above a pure liquid is shown on the right. The equilibrium vapor pressure is a result of the presence of molecules in the vapor phase above the molecules in the liquid phase. The figure below shows the affect on the vapor pressure of the solution upon addition of a nonvolatile solute.



where

 $\text{C}_{12}\text{H}_{22}\text{O}_{11}$

After addition of the solute

The addition of a nonvolatile solute decreases the vapor pressure due to the liquid in the vapor phase. This occurs because there is a decreased entropic drive for the vapor to form above the solution compared to the vapor forming above the pure liquid.

Section 18 3 Colligative Properties Of Solutions Answers

Marco Cascella



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