

# CHEMICAL REACTIONS

# **Practice Problems**

In your notebook, solve the following problems. Use the 3-step problem-solving approach you learned in Chapter I.

#### SECTION 11.1 DESCRIBING CHEMICAL REACTIONS

- Write the skeleton equation for the reaction between hydrogen and oxygen that produces water.
- Write the skeleton equation for the reaction that produces iron(II) sulfide from iron and sulfur.
- Write the skeleton equation representing the heating of magnesium carbonate to produce solid magnesium oxide and carbon dioxide gas.
- 4. Write a balanced equation for the production of HCl gas from its elements.
- Write a sentence that completely describes the chemical reaction represented by this balanced equation.

 $2HCl(aq) + CaCO_3(s) \rightarrow CO_2(g) + CaCl_2(aq) + H_2O(l)$ 

Write the word equation for the following equation. Write a sentence fully describing the reaction. Is the equation correctly balanced? Explain.

 $2Ag(s) + S(s) \rightarrow Ag_{s}S(s)$ 

- Write a balanced equation representing the formation of aqueous sulfuric acid from water and sulfur trioxide gas.
- 8. Write a balanced equation from this word equation.

aqueous silver nitrate + copper metal -- silver metal + aqueous copper nitrate

9. Write a balanced equation for the following word equation.

phosphorus + oxygen → tetraphosphorous decoxide

### SECTION 11.2 TYPES OF CHEMICAL REACTIONS

- Write a balanced equation representing the reaction of magnesium with oxygen gas to produce magnesium oxide.
- Write the balanced equation for the reaction that occurs between aluminum and fluorine.
- Write the balanced equation for the production of oxygen gas and potassium chloride from the decomposition of potassium chlorate.
- Write the balanced equation for the reaction between hydrochloric acid and calcium metal. The products are hydrogen gas and calcium chloride.
- Write the balanced equation for the combustion of propane (C<sub>3</sub>H<sub>8</sub>) to produce carbon dioxide and water vapor.
- Write the balanced equation for the reaction between iron(III) chloride and sodium hydroxide. The products are iron(III) hydroxide and sodium chloride.

# **Section 11 1 Describing Chemical Reactions Practice Problems Answers**

**Kathleen Armour** 

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## **Table of Contents Section 11 1 Describing Chemical Reactions Practice Problems Answers**

- 1. Understanding the eBook Section 11 1 Describing Chemical Reactions Practice Problems Answers
  - The Rise of Digital Reading Section 11 1 Describing Chemical Reactions Practice Problems Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Section 11 1 Describing Chemical Reactions Practice Problems Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Section 11 1 Describing Chemical Reactions Practice Problems Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Section 11 1 Describing Chemical Reactions Practice Problems Answers
  - Personalized Recommendations

#### Section 11 1 Describing Chemical Reactions Practice Problems Answers

- Section 11 1 Describing Chemical Reactions Practice Problems Answers User Reviews and Ratings
- Section 11 1 Describing Chemical Reactions Practice Problems Answers and Bestseller Lists
- 5. Accessing Section 11 1 Describing Chemical Reactions Practice Problems Answers Free and Paid eBooks
  - Section 11 1 Describing Chemical Reactions Practice Problems Answers Public Domain eBooks
  - Section 11 1 Describing Chemical Reactions Practice Problems Answers eBook Subscription Services
  - Section 11 1 Describing Chemical Reactions Practice Problems Answers Budget-Friendly Options
- 6. Navigating Section 11 1 Describing Chemical Reactions Practice Problems Answers eBook Formats
  - ∘ ePub, PDF, MOBI, and More
  - Section 11 1 Describing Chemical Reactions Practice Problems Answers Compatibility with Devices
  - Section 11 1 Describing Chemical Reactions Practice Problems Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - o Adjustable Fonts and Text Sizes of Section 11 1 Describing Chemical Reactions Practice Problems Answers
  - Highlighting and Note-Taking Section 11 1 Describing Chemical Reactions Practice Problems Answers
  - $\circ$  Interactive Elements Section 11 1 Describing Chemical Reactions Practice Problems Answers
- 8. Staying Engaged with Section 11 1 Describing Chemical Reactions Practice Problems Answers
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Section 11 1 Describing Chemical Reactions Practice Problems Answers
- 9. Balancing eBooks and Physical Books Section 11 1 Describing Chemical Reactions Practice Problems Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Section 11 1 Describing Chemical Reactions Practice Problems Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Section 11 1 Describing Chemical Reactions Practice Problems Answers
  - Setting Reading Goals Section 11 1 Describing Chemical Reactions Practice Problems Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section 11 1 Describing Chemical Reactions Practice Problems Answers
  - Fact-Checking eBook Content of Section 11 1 Describing Chemical Reactions Practice Problems Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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