Name Date Class

Chemical Reactions . Review and Reinforce

# **Describing Chemical Reactions**

#### **Understanding Main Ideas**

Balance the equations on the lines below. State whether the reaction is a synthesis, decomposition, or replacement reaction.

Given Equation	Balanced Equation	Type of Reaction
$f_* \text{ FeS} * \text{ HCL} \rightarrow \text{FeCl}_2 * \text{ H}_2\text{S}$	w.	ъ.
2. Na + F <sub>2</sub> → NaF	+.	b.
3. HgO → H <sub>S</sub> + O <sub>2</sub>	e.	ъ.

Answer questions 4 and 5 on a separate sheet of paper.

- Describe in words the chemical composition of the molecules involved and the reaction represented by the equation: 2H<sub>2</sub> + O<sub>2</sub> → 2 H<sub>2</sub>O
- Use the principle of conservation of mass to explain why the equation in question 6 is balanced.

#### **Building Vocabulary**

Match each term with its difficultion by writing the letter of the correct difficultion in the right column on the line beside the term in the lift column.

6. chemical equation	<ul> <li>a. substance present after a reaction.</li> </ul>
7. decomposition reaction	<ul> <li>b. reaction in which substances combine to form a more complex compound</li> </ul>
B. coefficient	<ul> <li>uses symbols and formulas to show cherefoll reactions</li> </ul>
9. product	<ul> <li>d. reaction in which one element replaces another in a compound.</li> </ul>
18. seactant 11. conservation of mass	e. substance present before a reaction
	<ol> <li>rrumber telling hose many molecules of a substance are involved in a chemical reaction.</li> </ol>
12. syrethesis maction	<ul> <li>p. reaction in which compounds are broken down into simpler products</li> </ul>
13- replacement reaction	<ul> <li>b. principle that states that matter is not created or destroyed during a chemical</li> </ul>

resoltion.

# **Section 1 2 Describing Chemical Reactions Review And Reinforce**

Joseph A. Mascetta, Mark Kernion

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