

5 • Chemical Bonding**LEWIS STRUCTURES**Indicate the # of VALENCE electrons for each species. Write the correct Lewis electron dot structure for each.

F # of electrons = <u>7</u> $\cdot \ddot{\text{F}} \cdot$	O # of electrons = <u>6</u> $\cdot \ddot{\text{O}} \cdot$ or $\cdot \ddot{\text{O}} \cdot$	K # of electrons = <u>1</u> $\text{K} \cdot$	Al # of electrons = <u>3</u> $\cdot \text{Al} \cdot$
F⁻ <u>7+1</u> # of electrons = <u>8</u> $[\cdot \ddot{\text{F}}:]^{-}$	O²⁻ <u>6+2</u> # of electrons = <u>8</u> $[\cdot \ddot{\text{O}}:]^{2-}$	K⁺ <u>1-1</u> # of electrons = <u>0</u> $[\text{K}]^{+}$	Al³⁺ <u>3-3</u> # of electrons = <u>0</u> $[\text{Al}]^{3+}$
F₂ <u>7+7</u> # of electrons = <u>14</u> $\cdot \ddot{\text{F}} : \ddot{\text{F}} \cdot$	H₂ <u>1+1</u> # of electrons = <u>2</u> $\text{H} : \text{H}$	HF <u>1+7</u> # of electrons = <u>8</u> $\text{H} : \ddot{\text{F}} \cdot$	NH₃ <u>5+3</u> # of electrons = <u>8</u> $\begin{array}{c} \text{H} : \ddot{\text{N}} : \text{H} \\ \\ \text{H} \end{array}$
CH₄ <u>4+4</u> # of electrons = <u>8</u> $\begin{array}{c} \text{H} \\ \\ \text{H} : \text{C} : \text{H} \\ \\ \text{H} \end{array}$	NF₃ <u>5+21</u> # of electrons = <u>26</u> $\begin{array}{c} \cdot \ddot{\text{F}} : \ddot{\text{N}} : \ddot{\text{F}} : \\ \\ \cdot \ddot{\text{F}} : \end{array}$	SiF₄ <u>4+28</u> # of electrons = <u>32</u> $\begin{array}{c} \cdot \ddot{\text{F}} : \\ \cdot \ddot{\text{F}} : \text{Si} : \ddot{\text{F}} : \\ \\ \cdot \ddot{\text{F}} : \end{array}$	C₂H₆ <u>8+6</u> # of electrons = <u>14</u> $\begin{array}{c} \text{H} \quad \text{H} \\ \quad \\ \text{H} : \text{C} : \text{C} : \text{H} \\ \quad \\ \text{H} \quad \text{H} \end{array}$
MgH₂ <u>2+2</u> # of electrons = <u>4</u> $\text{H} : \text{Mg} : \text{H}$	LiH <u>1+1</u> # of electrons = <u>2</u> $\text{Li} : \text{H}$	AlH₃ <u>3+3</u> # of electrons = <u>6</u> $\begin{array}{c} \text{H} : \text{Al} : \text{H} \\ \\ \text{H} \end{array}$	BH₃ <u>3+3</u> # of electrons = <u>6</u> $\begin{array}{c} \text{H} : \text{B} : \text{H} \\ \\ \text{H} \end{array}$

Practice Sheet Electron Dot Lewis Structures

CH Cherryholmes



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postgraduate researchers The rigorous yet accessible treatment of each subject area is ideal for those wanting a primer in a given topic to prepare them for more advanced study or research The learning features provided including questions at the end of every chapter and online multiple choice questions encourage active learning and promote understanding Furthermore frequent diagrams margin notes and glossary definitions all help to enhance a student s understanding of these essential areas of chemistry Chemical bonding gives a clear and succinct explanation of this fundamental topic which underlies the structure and reactivity of all molecules and therefore the subject of chemistry itself Little prior knowledge or mathematical ability is assumed making this the perfect text to introduce students to the subject

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