



Name _____ Class _____ Date _____

1

How many atoms of **oxygen (O)** are in the **chemical formula** below?

- A 4
- B 5
- C 7
- D 12

**3**

Carbon (C) and hydrogen (H) create **covalent bonds** to share **electrons** between them. What is the total number of electrons being shared between **carbon** and **hydrogen** in the molecule of propane shown below?

- A 8
- B 11
- C 16
- D 32

**2**

The diagram shown below is a **structural formula** for the gas **propane**. What would be the correct **chemical formula** for propane?

- A C_3H_6
- B CH_6
- C $8\text{H}_3\text{C}$
- D CH

**4**

The **reason** why the bonds between carbon (C) and hydrogen (H) are **covalent** is _____

- A C and H do not have electrons to give away
- B C and H are both nonmetals
- C C and H are both metalloids
- D C and H can bond only by sharing electrons

**5**

The diagram below represents the **structural formula** for several **water molecules**. How many **molecules** are represented?

- A 3
- B 5
- C 10
- D 15

**6**

Explain the **arrangement** of the molecules in the diagram of water below.

- A it is a lucky arrangement
- B water molecules are sticky
- C water molecules have polarity
- D water molecules are magnetic

**7**

Water molecules have some **polarity** in their structure. Using the diagram below, notice how the ends of H atoms (**positive**) and O atoms (**negative**) meet. This property is the **reason** that _____

- A water forms droplets easily
- B water boils easily
- C ice melts easily
- D water has color

**8**

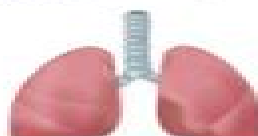
Since H atoms are slightly **positive** and the O atoms are slightly **negative**, molecules of water have **polarity**. This makes water a **good** _____

- A solute
- B refrigerant
- C solvent
- D magnet

**9**

The chemical formula for the **oxygen** that we breathe is _____

- A O
- B O_2
- C O_3
- D 2O

**10**

How many **molecules of oxygen** are represented in the chemical formula below?

- A 2
- B 4
- C 6
- D 18



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A Loxley



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web apr 10 2020 chemistry sk025 so this is my next upload on this web page will only contain of chemistry notes exercise and exam sets it is to be said that semester 2 chemistry will much harder than semester 1 chemistry in total there is 13 chapter in this semester 2 the first 3 chapters about physical chemistry while the rest are organic

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