

Calculus-Based Properties (Practice Problems)

1. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of glucose, $C_6H_{12}O_6$, and 100 g of water? Equilibrium is at 100°C and $P_g = 1$ atm.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{180 \text{ g}}}{0.1 \text{ kg}} = 5.56 \text{ mol/kg}$$

$$\Delta T_f = \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1.86^\circ\text{C}}{\text{mol}} \times 5.56 \text{ mol/kg} = 10.4^\circ\text{C}$$

$\Delta T_f = i K_f m$
 $\Delta T_f = 1 \times 1.86^\circ\text{C} \times 5.56 \text{ mol/kg}$
 $\Delta T_f = 10.4^\circ\text{C}$
 $T_f = 0^\circ\text{C} - 10.4^\circ\text{C} = -10.4^\circ\text{C}$
 $T_b = 100^\circ\text{C} + 10.4^\circ\text{C} = 110.4^\circ\text{C}$

2. What will be the freezing-point and boiling-point of an aqueous solution containing 100 g of urea, CH_4N_2O , and 100 g of water? Equilibrium is at 100°C and $P_g = 1$ atm. Assume $\Delta H_{\text{vap}} = 40.7 \text{ kJ/mol}$.

$$\text{Molality} = \frac{\text{moles solute}}{\text{kg solvent}} = \frac{100 \text{ g} \times \frac{1 \text{ mol}}{60 \text{ g}}}{0.1 \text{ kg}} = 16.7 \text{ mol/kg}$$

$$\Delta T_f = \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1.86^\circ\text{C}}{\text{mol}} \times 16.7 \text{ mol/kg} = 31.1^\circ\text{C}$$

$\Delta T_f = i K_f m$
 $\Delta T_f = 1 \times 1.86^\circ\text{C} \times 16.7 \text{ mol/kg}$
 $\Delta T_f = 31.1^\circ\text{C}$
 $T_f = 0^\circ\text{C} - 31.1^\circ\text{C} = -31.1^\circ\text{C}$
 $T_b = 100^\circ\text{C} + 31.1^\circ\text{C} = 131.1^\circ\text{C}$

3. A 100 g mass of urea, CH_4N_2O , is dissolved in 100 g of water. Assuming the boiling-point of the solution is 105.5°C, assume that the van't Hoff factor is 1.00. Find the molality of the solution.

$$\Delta T_b = T_b - T_f = 105.5^\circ\text{C} - 100^\circ\text{C} = 5.5^\circ\text{C}$$

$$\Delta T_b = i K_b m$$

$$5.5^\circ\text{C} = 1 \times 0.51^\circ\text{C} \times m$$

$$m = 10.8 \text{ mol/kg}$$

$$m = \frac{\text{moles solute}}{\text{kg solvent}}$$

$$10.8 \text{ mol/kg} = \frac{\text{moles solute}}{0.1 \text{ kg}}$$

$$\text{moles solute} = 1.08 \text{ mol}$$

$$\text{moles urea} = \frac{\text{mass urea}}{\text{molar mass urea}} = \frac{100 \text{ g}}{60 \text{ g/mol}} = 1.67 \text{ mol}$$

4. A solution composed from 10 g of urea, CH_4N_2O , and 100 g of water is composed. The freezing-point of this solution is 0°C. The normal freezing point of pure water is 0°C and the freezing-point depression constant for water is 1.86°C/kg . Determine the molar mass of the compound. Assume $\Delta H_{\text{vap}} = 40.7 \text{ kJ/mol}$.

$$\Delta T_f = \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1.86^\circ\text{C}}{\text{mol}} \times m = 0^\circ\text{C} - 0^\circ\text{C} = 0^\circ\text{C}$$

$$0^\circ\text{C} = 1.86^\circ\text{C} \times m$$

$$m = 0 \text{ mol/kg}$$

$$m = \frac{\text{moles urea}}{\text{kg water}}$$

$$m = \frac{\text{moles urea}}{\text{kg water}}$$

$$0 \text{ mol/kg} = \frac{\text{moles urea}}{0.1 \text{ kg}}$$

$$\text{moles urea} = 0 \text{ mol}$$

$$\text{moles urea} = \frac{\text{mass urea}}{\text{molar mass urea}} = \frac{10 \text{ g}}{M}$$

$$0 \text{ mol} = \frac{10 \text{ g}}{M} \Rightarrow M = \frac{10 \text{ g}}{0} = \infty$$

Rosengarten Colligative Properties Homework Answers

S Marginson



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