

Spring Chapter 6

1. Consider the following energy level diagram for a reaction.

(Energy)

at reactants is 100 kJ

(kJ/mol)

at products is 200 kJ

(kJ/mol)

2. Which of the following is the correct energy diagram for the reaction?



3. A reaction has a single peak. The reactants are at a higher energy level than the products. The energy axis is labeled (Energy) and (kJ/mol). Which of the following is the correct energy diagram for the reaction?

• $\Delta H = -100 \text{ kJ/mol}$
The reaction is exothermic.

• $\Delta H = +100 \text{ kJ/mol}$
The reaction is endothermic.

- Both products
- Reactants
- Both
- None

4. Which of the following is the correct energy diagram for the reaction?

5. Consider the reaction of hydrogen and oxygen to form water. The reaction is exothermic. Which of the following is the correct energy diagram for the reaction?

• $\Delta H = -286 \text{ kJ/mol}$
The reaction is exothermic.

• $\Delta H = +286 \text{ kJ/mol}$
The reaction is endothermic.

6. 21.40 g of a sample of a substance is heated from 25.0°C to 35.0°C. The heat of fusion is 10.0 kJ/mol. Which of the following is the correct energy diagram for the reaction?

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