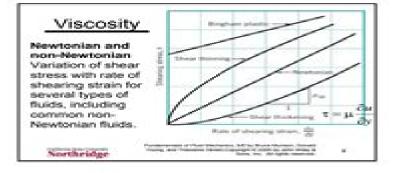
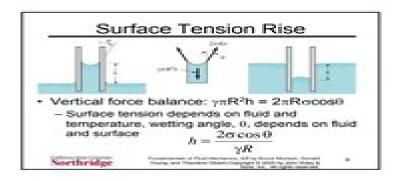
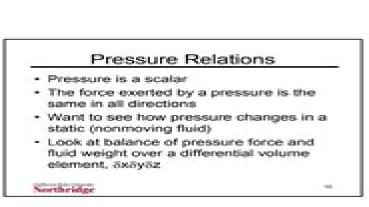
#### Ideal Gases

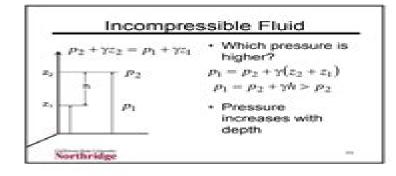
- From chemistry: PV = n9RT (V is volume)
  - n = m / M is the number of moles
    - for mass in kg n is in kilogram moles (kmol); for mass in lb<sub>m</sub>, n is in pound moles (lbmol)
  - R = 8.31447 k,J/kmol-K = 10.7316 psia-ft<sup>3</sup> / Ibmol-R is universal gas constant
  - R = PMM is engineering gas constant that is different for each gas
  - Real gases like ideal gases at low pressures
  - $-P = n\Re T / V = (m/M)\Re T / V = (m/V)(\Re M)T$

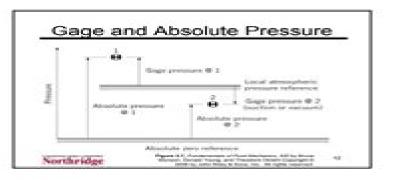
Northridge  $P = \rho RT$ 











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