

7-1 Review and Reinforcement

Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative
positive
crisscross

noble gas
electrons
anion

octet
ions
cation

electron configuration
polyatomic
empirical formula

1. An ionic bond is an attraction between oppositely charged ions.
2. Anions have a negative charge.
3. An atom becomes an ion by losing or gaining electrons.
4. The octet rule states that atoms tend to gain, lose, or share electrons in order to acquire a full set of valence electrons.
5. When sodium and chlorine form an ionic bond, both ions acquire the electron configuration of a(n) noble gas.
6. The atoms that make up polyatomic ions are bonded together by covalent bonds.
7. The crisscross method can be used to write the formula for an ionic compound.
8. The empirical formula of a compound denotes the ratio of ions in the compound.

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line.

losing

true

true

polyatomic

seven

true (released)

9. Calcium becomes a monatomic cation by gaining two electrons.

10. A cation has a positive charge.

11. A binary ionic compound contains only one kind of cation and one kind of anion.

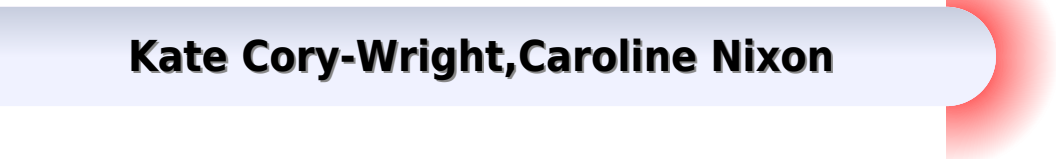
12. Monatomic ions consist of more than one atom.

13. The Lewis dot diagram for chlorine, a group 7A element, has six electrons.

14. A great deal of energy is produced when an ionic compound is formed.

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Kate Cory-Wright, Caroline Nixon



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Table of Contents Review And Reinforcement Worksheet Answers

1. Understanding the eBook Review And Reinforcement Worksheet Answers
 - The Rise of Digital Reading Review And Reinforcement Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Review And Reinforcement Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review And Reinforcement Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review And Reinforcement Worksheet Answers
 - Personalized Recommendations
 - Review And Reinforcement Worksheet Answers User Reviews and Ratings
 - Review And Reinforcement Worksheet Answers and Bestseller Lists
5. Accessing Review And Reinforcement Worksheet Answers Free and Paid eBooks
 - Review And Reinforcement Worksheet Answers Public Domain eBooks
 - Review And Reinforcement Worksheet Answers eBook Subscription Services
 - Review And Reinforcement Worksheet Answers Budget-Friendly Options
6. Navigating Review And Reinforcement Worksheet Answers eBook Formats

- ePub, PDF, MOBI, and More
- Review And Reinforcement Worksheet Answers Compatibility with Devices
- Review And Reinforcement Worksheet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Review And Reinforcement Worksheet Answers
 - Highlighting and Note-Taking Review And Reinforcement Worksheet Answers
 - Interactive Elements Review And Reinforcement Worksheet Answers
- 8. Staying Engaged with Review And Reinforcement Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review And Reinforcement Worksheet Answers
- 9. Balancing eBooks and Physical Books Review And Reinforcement Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review And Reinforcement Worksheet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Review And Reinforcement Worksheet Answers
 - Setting Reading Goals Review And Reinforcement Worksheet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Review And Reinforcement Worksheet Answers
 - Fact-Checking eBook Content of Review And Reinforcement Worksheet Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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