Stoichiometry Review

The combustion of propane occurs via the reaction:
 C₃H₀(g) + 5 O₂(g) → 3 CO₂(g) + 4 H₂O(g)
 How many grams of oxygen are required to burn completely 10.0 g of propane?

Small quantities of chlorine can be prepared in the laboratory by the reaction:
 MnO₂(s) + 4 HCl(aq) → MnCl₂(aq) + Cl₂(g) + 2 H₂O(l)
 What volume of chlorine at STP can be prepared from 100.0 g of manganese dioxide?

The most common are of arsenic is mispickel, Fe5As. Upon heating this are, free arsenic is obtained. How many grams of Fe5As are required to produce 10.0 g of arsenic according to the following reaction?
 Fe5As(s) → Fe5(s) + As(s)

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J Rink

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