

AP® CHEMISTRY 2005 EXAMINATIONS CHEMISTRY SECTION

Question 10
(8 points)

As sample of $\text{C}_2\text{H}_2\text{Cl}_2$ is subjected to electrolysis, it is found that 0.010 moles of chlorine gas are formed for every 0.0050 moles of the sample. An electrochemical cell is constructed as follows:



Use the standard E° in brackets for the reduction half-reaction. The temperature of the $\text{C}_2\text{H}_2\text{Cl}_2$ is 25°C. (Assume that the activity of the ions is unity.)

Answer each part:

- Write the half-reaction occurring at the cathode in the electrochemical cell. Indicate whether the reaction is spontaneous or nonspontaneous.
- Write the half-reaction occurring at the anode in the electrochemical cell. Indicate whether the reaction is spontaneous or nonspontaneous.

1 point is earned for each half-reaction.

$\text{C}_2\text{H}_2\text{Cl}_2$ reacts readily with H_2O to produce $\text{C}_2\text{H}_4\text{Cl}_2$, as represented by the following equation:



The $\text{C}_2\text{H}_2\text{Cl}_2$ is a liquid at room temperature, whereas $\text{C}_2\text{H}_4\text{Cl}_2$ is a liquid at room temperature. Ethane is the only hydrocarbon to boil at a temperature at which the standard enthalpy of formation of the hydrocarbon is zero at 25°C.

Use standard enthalpies of formation to calculate the standard enthalpy of formation of $\text{C}_2\text{H}_2\text{Cl}_2$. Assume that the standard enthalpy of formation of $\text{C}_2\text{H}_4\text{Cl}_2$ is zero.

1 point is earned for the calculation.

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CL Gary



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