

AP® CHEMISTRY 2005 EXAMINATIONS CHEMISTRY SECTION

Question 10
(8 points)

As sample of $\text{C}_2\text{H}_2\text{Br}_4$ is subjected to continuous vacuum, liquid $\text{C}_2\text{H}_2\text{Br}_4$ condenses and then boils at 2.0 atm. The pressure of the sample is continuously reduced throughout the experiment.



Use the graph to determine the temperature of liquid $\text{C}_2\text{H}_2\text{Br}_4$ at a pressure of 1.0 atm. Assume that the pressure of the sample is reduced slowly.

Answer the two:

- (a) as the temperature decreases, the temperature of the condensed $\text{C}_2\text{H}_2\text{Br}_4$ increases, the temperature of the condensed $\text{C}_2\text{H}_2\text{Br}_4$ decreases, the temperature of the condensed $\text{C}_2\text{H}_2\text{Br}_4$ remains constant.
- (b) as the temperature decreases, the temperature of the condensed $\text{C}_2\text{H}_2\text{Br}_4$ increases, the temperature of the condensed $\text{C}_2\text{H}_2\text{Br}_4$ decreases, the temperature of the condensed $\text{C}_2\text{H}_2\text{Br}_4$ remains constant.

1 point is earned for each correct answer.

$\text{C}_2\text{H}_2\text{Br}_4$ reacts readily with H_2O to produce $\text{C}_2\text{H}_2\text{Br}_2$ and H_2OBr_2 as represented by the following equation.



One of the $\text{C}_2\text{H}_2\text{Br}_4$ liquid-saturated vessels contained with $\text{C}_2\text{H}_2\text{Br}_4$ at 2.0 atm, contained $\text{C}_2\text{H}_2\text{Br}_4$ and H_2O . The vessel was cooled to 100 K, resulting in condensation of $\text{C}_2\text{H}_2\text{Br}_4$ and H_2O and formation of $\text{C}_2\text{H}_2\text{Br}_2$ and H_2OBr_2 .

Use the data in the table below to determine the temperature of the condensed $\text{C}_2\text{H}_2\text{Br}_4$ and H_2O at 100 K. Assume that the pressure of the condensed $\text{C}_2\text{H}_2\text{Br}_4$ and H_2O is 1.0 atm.

1 point is earned for each correct answer.

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**Amaya Azqueta, Sabine Langie, Andrew
Collins**



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