

Name Key Class _____ Date _____

Nuclear Decay

Types of Radiation

1. Complete the table below to compare the properties of alpha, beta, positron, and gamma radiation.

Name	Alpha	Beta	Positron	Gamma
Greek letter	α	β^-	β^+	γ
Symbol	${}^4_2\text{He}$	${}^0_{-1}e$	${}^0_{+1}e$	N/A
Composition	$2p^+$, $2n^0$	electron	anti-electron	EM wave
Charge	$2+$	$1-$	$1+$	0
Stopped by	paper, skin, clothing	wood, glass	wood, glass	thick concrete or lead

Nuclear Decay Reactions

In problems 2 – 5, fill in the missing particle for each nuclear decay reaction. Then, match each equation to the appropriate type of nuclear decay.



A. alpha emission



B. beta emission



C. electron capture



D. positron emission

In problems 6 and 7, write a balanced equation for each nuclear decay reaction.

6. Decay of radium-223 by alpha (α) emission.



7. Decay of iodine-126 by positron (β^+) emission.



8. How are the atomic number and mass number of a nuclide affected by positron emission?

Atomic number decreases by one. Mass number stays the same.

9. How are the atomic number and mass number of a nuclide affected by the emission of gamma radiation?

Both atomic number and mass number stay the same since gamma radiation has no mass or charge.

10. Which has a slightly greater mass in a nuclear reaction, the reactants or the products? Why?

The reactants have a slightly greater mass. In a nuclear reaction, a small amount of mass is converted to energy according to the equation $E = mc^2$. The difference in mass is referred to as the mass defect.

Radioactive Decay Study Guide Answer Key

Tao Wei



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Table of Contents Radioactive Decay Study Guide Answer Key

1. Understanding the eBook Radioactive Decay Study Guide Answer Key
 - The Rise of Digital Reading Radioactive Decay Study Guide Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Radioactive Decay Study Guide Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Radioactive Decay Study Guide Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Radioactive Decay Study Guide Answer Key
 - Personalized Recommendations
 - Radioactive Decay Study Guide Answer Key User Reviews and Ratings
 - Radioactive Decay Study Guide Answer Key and Bestseller Lists
5. Accessing Radioactive Decay Study Guide Answer Key Free and Paid eBooks
 - Radioactive Decay Study Guide Answer Key Public Domain eBooks
 - Radioactive Decay Study Guide Answer Key eBook Subscription Services
 - Radioactive Decay Study Guide Answer Key Budget-Friendly Options

6. Navigating Radioactive Decay Study Guide Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Radioactive Decay Study Guide Answer Key Compatibility with Devices
 - Radioactive Decay Study Guide Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Radioactive Decay Study Guide Answer Key
 - Highlighting and Note-Taking Radioactive Decay Study Guide Answer Key
 - Interactive Elements Radioactive Decay Study Guide Answer Key
8. Staying Engaged with Radioactive Decay Study Guide Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Radioactive Decay Study Guide Answer Key
9. Balancing eBooks and Physical Books Radioactive Decay Study Guide Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Radioactive Decay Study Guide Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Radioactive Decay Study Guide Answer Key
 - Setting Reading Goals Radioactive Decay Study Guide Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Radioactive Decay Study Guide Answer Key
 - Fact-Checking eBook Content of Radioactive Decay Study Guide Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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