

14

THE BEHAVIOR OF GASES

Chapter Test II

A. Matching

Match each term in Column I with the correct description in Column II. Write the letter of the correct description next to the term.

Column I

1. The constant relating unit temperature changes to changes in volume for a fixed amount of gas at constant pressure (the ratio of the change in volume to the change in temperature).
2. The volume of a fixed amount of a gas at constant pressure in the standard state (0°C and 1 atm).
3. The ratio of the mass of a gas to its volume, independent of the actual amount of gas used.
4. An experimental gas constant determined from the ideal gas law at 0°C and 1 atm.
5. A constant of nature that has values of about 8.314 J/mol·K.
6. The a parameter of gas constant equations for determining the gas molecular weight of a gas.
7. An experimental gas constant derived from the ideal gas law under standard conditions (0°C and 1 atm).
8. $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$
9. A constant of nature that has a value of 2.303.
10. A constant of nature that has a value of 10.

Column II

- a. standard gas law
- b. ideal gas constant R
- c. density
- d. compressibility
- e. Boyle's law
- f. molar volume
- g. standard molar volume
- h. volume
- i. compressibility
- j. compressibility coefficient

B. Multiple Choice

Choose the correct answer and write its letter on the line.

1. According to Graham's law, the rate of effusion of a gas is proportional to the square root of its molecular weight. Which of the following gases effuses most rapidly?
a. the heaviest gas
b. the lightest gas
c. the gas with the highest molecular weight
d. the gas with the lowest molecular weight

Prentice Hall Chemistry 8 Test

Mervyn Richardson



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1. Understanding the eBook Prentice Hall Chemistry 8 Test
 - The Rise of Digital Reading Prentice Hall Chemistry 8 Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Prentice Hall Chemistry 8 Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Prentice Hall Chemistry 8 Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Prentice Hall Chemistry 8 Test
 - Personalized Recommendations
 - Prentice Hall Chemistry 8 Test User Reviews and Ratings
 - Prentice Hall Chemistry 8 Test and Bestseller Lists
5. Accessing Prentice Hall Chemistry 8 Test Free and Paid eBooks
 - Prentice Hall Chemistry 8 Test Public Domain eBooks
 - Prentice Hall Chemistry 8 Test eBook Subscription Services
 - Prentice Hall Chemistry 8 Test Budget-Friendly Options

6. Navigating Prentice Hall Chemistry 8 Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Prentice Hall Chemistry 8 Test Compatibility with Devices
 - Prentice Hall Chemistry 8 Test Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Prentice Hall Chemistry 8 Test
 - Highlighting and Note-Taking Prentice Hall Chemistry 8 Test
 - Interactive Elements Prentice Hall Chemistry 8 Test
8. Staying Engaged with Prentice Hall Chemistry 8 Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Prentice Hall Chemistry 8 Test
9. Balancing eBooks and Physical Books Prentice Hall Chemistry 8 Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Prentice Hall Chemistry 8 Test
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Prentice Hall Chemistry 8 Test
 - Setting Reading Goals Prentice Hall Chemistry 8 Test
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Prentice Hall Chemistry 8 Test
 - Fact-Checking eBook Content of Prentice Hall Chemistry 8 Test
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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