

Question	Expected Answers	Marks	Additional Guidance
3 a	<p>measured pH &gt; 1 OR <math>[H^+] &lt; 0.1 \text{ (mol dm}^{-3}\text{)}</math> ✓</p> <p><math>[H^+] = 10^{-0.4}</math> ✓</p> <p><math>K_a = \frac{[H^+][CH_3CH_2COO^-]}{[CH_3CH_2COOH]}</math> OR <math>\frac{[H^+]^2}{[CH_3CH_2COOH]}</math> ✓</p> <p>Calculate <math>K_a</math> from <math>\frac{[H^+]^2}{0.100}</math> ✓</p>	4	<p>ALLOW <math>C_2H_3</math> throughout question</p> <p>ALLOW <math>[H^+] &lt; [CH_3CH_2COOH]</math> OR <math>[H^+] &lt; [HA]</math>  ALLOW measured pH is higher than expected  ALLOW measured pH is not as acidic as expected  ALLOW a quoted pH value or range &gt; 1 and &lt; 7  OR between 1 and 7</p> <p>ALLOW <math>[H^+] = \text{antilog } -\text{pH}</math> OR <math>[H^+] = \text{inverse log } -\text{pH}</math></p> <p>ALLOW <math>\frac{[H^+][A^-]}{[HA]}</math> OR <math>\frac{[H^+]^2}{[HA]}</math></p> <p>IF <math>K_a</math> is NOT given and <math>K_a = \frac{[H^+]^2}{0.100}</math> is shown, award mark for <math>K_a</math> also  (I.e. <math>K_a = \frac{[H^+]^2}{0.100}</math> is automatically awarded the last 2 marks)</p>
b	<p>Marks are for correctly calculated values.  Working shows how values have been derived.</p> <p><math>[H^+] = 10^{-13.46} = 3.47 \times 10^{-14} \text{ (mol dm}^{-3}\text{)}</math> ✓</p> <p><math>[OH^-] = \frac{1.0 \times 10^{-14}}{3.47 \times 10^{-14}} = 0.29 \text{ (mol dm}^{-3}\text{)}</math> ✓</p>	2	<p>ALLOW <math>3.467368505 \times 10^{-14}</math> and correct rounding to <math>3.5 \times 10^{-14}</math></p> <p>ALLOW 0.28840315 and correct rounding to 0.29,  i.e. ALLOW 0.288</p> <p>ALLOW alternative approach using pOH:</p> <p>pOH = <math>14 - 13.46 = 0.54</math> ✓  <math>[OH^-] = 10^{-0.54} = 0.29 \text{ (mol dm}^{-3}\text{)}</math> ✓</p> <p>Correct answer gets BOTH marks</p>

# Ocr F325 June 2010 Question Paper

**Albert A Gayle**



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