		9
d)	In a	an experiment, a student makes a solution of strontium chloride, $SrCl_2$ , by adding excess the hydrochloric acid to strontium carbonate.
	(i)	Describe what the student would observe and write the equation for the reaction.
		observations
		equation[2]
	(ii)	Draw a 'dot-and-cross' diagram to show the bonding of strontium chloride. Show <b>outer</b> electrons only.
		[2]
e)	In another experiment, a student attempts to make a solution of strontium chloride by adding chlorine water to aqueous strontium bromide.	
	(i)	Describe what the student would observe.
		[1]
	(ii)	Write the ionic equation for the reaction which takes place.
		[1]
	(iii)	Chlorine is more reactive than bromine. Explain why.
		[4]
		[Total: 18]

Turn over

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