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Oce 7056 User Manual

M Carnoy

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web objective to apply stoichiometric principles to calculate the appropriate chemical equation for the reaction between iron
metal and a solution of copper i sulfate introduction in the form of finely split powder copper metal is produced as a
byproduct of the process as iron ions the iron metal dissolves

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of 8 001 grams of copper ii sulfate CuSO_4 and 2 0153 grams of iron powder Fe the goal of this experiment was to determine
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web procedure 1 find the mass of a clean dry 100 ml beaker and record 2 measure out 8 00 grams of copper ii sulfate onto a

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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away cooper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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they were happy enough it is safer to search in the maze than remain in a cheeseless situation

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