2 A.M.U. Sci./Dip. Engg.		2007-2008
(c) 350 cm, convergent		(c) Cl ₂ and CaSO ₄ (d) Cl ₂ and Ca(HCO ₃) ₂
	27	The hard glass is obtained by fusing:
Three resistances A. B and C each of 50 and	-	A to the second of the second second black and the second
connected to a battery of 10 V as about to the		(b) a mixture of sand, lime, borax and arkan
		(a) autosium carbonate and limescone
W		(4) potentium early-posts, lead oxide and sand
#v [⊥]	28.	Which of the following metals is most reactive?
		(a) Aluminium (b) lead
() 0		(c) Mercury (d) Silver Which of the following alloys contains
	29.	
		chromium? (a) Steel (b) Stainless steel
ine device used to generate electrical energy		
	-	(c) Magnalium (d) Brass Which of the following is monomer of
	30.	natural rubber?
		(a) Chloroethene (b) Chloroprene
(a) 0.25 X 10 ²³ (b) 0.50 X 10 ²³		(c) Isoprene (d) Buta-1, 3-diene
		Alkaline KMnO ₄ oxidises propanone to:
The value of charge / mass ratio of electron	31.	(a) propanoic acid (b) ethanoic acid
		(c) methanoic acid (d) oxalic acid
(a) W.K. Roenteen (b) I.I. Thomson	22	Which of the following metals can displace Zn
	34.	from ZnSO ₄ solution?
		(a) Calcium (b) Copper
example of combination reaction?		(c) Iron (d) Mercury
(a) $Zn(s) + CuSO_{s}(sa) \rightarrow ZnSO_{s}(sa) + Cu(s)$	2.2	Heating of sodium ethanoate with soda lime
(b) BaCl ₂ (aq) + Na ₂ SO ₂ (aq) → BaSO ₂ (s) + 2NaCl(aq)	33.	vields:
(c) $CaCO_3(s) \rightarrow CaO(s) + CO_2(g)$		(a) ethane (b) methane
		(c) ethanol (d) methanol
	2.4	The compound formed by the reaction of
	34.	ethyne with bromine is:
		(a) Br-CH=CH-Br (b) Br-CH ₂ -CHBr ₂
The state of the s		(c) CH ₂ =CH-Br (d) Br ₂ CH-CHBr ₂
	2.6	Cell were first discovered by Robert Hooke in
	i altaba	the year:
		(a) 1665 (b) 1674 (c) 1831 (d) 1835
	24	The plant cells have a rigid cell wall that lies:
	30.	(a) Outside the plasma membrane
(c) Li <na<k<rb (d)="" rb<k<na<li<="" td=""><td></td><td>A CAN TO THE PROPERTY OF THE P</td></na<k<rb>		A CAN TO THE PROPERTY OF THE P
the state of the s		(b) Inside the plasma membrane
		(c) In between the plasma membranes
(a) Hydrogen (H) (b) Lithium (Li)	1900	(d) None of the above
(c) Oxygen (O) (d) Chlorine (Cl)	37.	During mitosis nucleolus and nuclea
Which of the following is an example of		membrane are lost in stage:
strong electrolyte?		(a) Prophase (b) Metaphase
(a) H ₂ CO ₁ (b) NH ₂ OH	25.00	(c) Anaphase (d) Telophase
(c) NaCl (d) COOH - COOH	38.	An example of simple and permanent tissue is
	1, 1414	(a) Xylem (b) Phloem
		(c) Selerenchyma (d) All the above
	39.	The shape of squamous epithelial tissue is:
		(a) Cubical (b) Flattened
		(c) Pillar like (d) None of the above
	40.	Fresh water sponge belongs to phylum:
	+	(a) Arthropoda (b) Annelida
(a) CaCi2 and CaCO3 (b) Ci2 and Ca(On)2		(o) rumenda
	Three resistances A, B and C each of 5Ω are connected to a battery of 10 V as shown in the figure. Calculate the current through C: SO	(c) 350 cm, convergent (d) 150 cm, divergent Three resistances A, B and C each of 5Ω are connected to a battery of 10 V as shown in the figure. Calculate the current through C: 28. (a) 2 amp (b) 1 amp 29. (c) 3 amp (d) 0.6 amp The device used to generate electrical energy is: (a) Electric motor (b) Generator (c) Galvanometer (d) Voltmeter The number of molecules in 11 g of CO ₂ are: (a) 0.25 X 10 ²⁹ (b) 0.50 X 10 ²⁹ (c) 1.00 X 10 ²⁹ (d) 1.51 X 10 ²⁹ 31. The value of charge / mass ratio of electron was determined by: (a) W K Roentgen (b) J.J.Thomson (c) Marie Curie (d) Niels Bohr Which of the following reactions is an example of combination reaction? (a) 2a(t) + Cu5O ₂ (a(t) + 2a5O ₂ (a(t) + Cu(t)) (b) 8c(3(t) + 3a5O ₂ (a(t) + BaSO ₂ (t) + 2hsC(t)a(t) (c) CaCO ₂ (t) → CO ₂ (t) (d) C(s) + Cy(g) → CO ₂ (t) The element 'X' and 'Y' have atomic numbers 12 and 17 respectively. Element 'X' reacts with element 'Y' to form compound with molecular formula: (a) XY (b) XY ₂ (c) XY ₃ (d) Xyy (b) XY ₂ (c) XY ₃ (d) Xyy (d) Xyy (e) XY ₃ The atomic radius (pm) of Li, Na, K and Rb varies in the order: (a) Na <k<rb<1 (b)="" k<na<li="">Rb<(S) (d) Rb<k<na<li (a)="" (kj="" (pl)="" affinity="" electron="" fluorine="" hyco<sub="" is="" less="" mol¹)="" of="" than:="" the="" value="">2 (b) NH₄OH (c) NaCl (d) CoOH – COOH If pH of a solution changes from 5 to 4, the change in hydrogen ion concentration shall be: (a) two times (b) five times (c) ten times (d) twenty times Bleaching powder is manufactured by the</k<na</k<rb<1>

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