

Chemistry 1020 Final Exam Review

1. A 25.0 mL sample of 1.50 M sodium chlorate (NaClO) is titrated with 1.00 M HCl . The K_a of HClO is 2.9×10^{-8} .

Weak Acid

(a) What is the pH after 20.0 mL of HCl has been added?



$n_{\text{ClO}^-} = (1.50\text{M})(0.0250\text{L}) = 0.0375\text{mol}$ ClO^- is in excess

$n_{\text{H}_3\text{O}^+} = (1.00\text{M})(0.0200\text{L}) = 0.0200\text{mol}$ all H_3O^+ reacts

buffer: still be on between start and equivalence point both ClO^- and HClO are in solution

$\text{pH} = \text{pKa} + \log \frac{[\text{ClO}^-]}{[\text{HClO}]}$

$\text{pKa} = \text{p}K_a = -\log K_a = -\log(2.9 \times 10^{-8}) = 7.49$

$n_{\text{ClO}^-} \text{ remaining} = n_{\text{ClO}^-} - n_{\text{H}_3\text{O}^+} = 0.0375\text{mol} - 0.0200\text{mol} = 0.0175\text{mol}$
 $n_{\text{HClO}} = n_{\text{H}_3\text{O}^+} = 0.0200\text{mol}$
 $\text{pH} = 7.49 + \log \left(\frac{0.0175}{0.0200} \right) = 7.44$

(b) What is the pH at the equivalence point?

eq. point: an ClO^- consumed. n_{HClO} added = n_{ClO^-} from 3rd



$V_{\text{eq}} = n_{\text{ClO}^-} / C_{\text{HCl}}$

$= \frac{0.0375\text{mol}}{1.00\text{M}} = 0.0375\text{L}$

$[\text{HClO}] = \frac{n_{\text{HClO}}}{V_{\text{total}}} = \frac{0.0375\text{mol}}{0.025\text{L} + 0.0375\text{L}} = 0.600\text{M}$

$\text{pH} = -\log(0.600) = -\log 0.600 = 0.22$

(c) What is the pH after 40.0 mL of HCl have been added?

eq. V of HCl was 37.5 mL

V of HCl after eq. point = $40.0\text{mL} - 37.5\text{mL} = 2.5\text{mL}$
 $= 0.0025\text{L}$

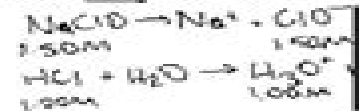
$n_{\text{H}_3\text{O}^+} \text{ in solution} = (0.0025\text{L})(1.00\text{M}) = 0.0025\text{mol} = n_{\text{H}_3\text{O}^+}$

$[\text{H}_3\text{O}^+] = \frac{n_{\text{H}_3\text{O}^+}}{V_{\text{total}}} = \frac{0.0025\text{mol}}{0.025\text{L}} = 0.100\text{M}$

$\text{pH} = -\log(0.100) = -\log(0.100) = 1.00$

Weak Acid

Strong Acid



C-17

Review Problems For Regular Chemistry Final

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