

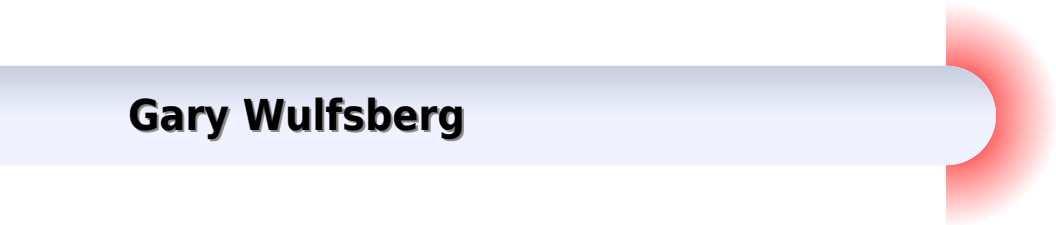
- Given the nuclear equation:
 $^{14}_7\text{N} + X \rightarrow ^{16}_8\text{O} + ^1_1\text{H}$
 What is particle X ?
 A) an alpha particle B) a beta particle
 C) a deuteron D) a triton
- The nucleus of a radium-226 atom is unstable, which causes the nucleus to spontaneously
 A) absorb electrons B) absorb protons
 C) decay D) oxidize
- Which group of nuclear emissions is listed in order of increasing charge?
 A) alpha particle, beta particle, gamma radiation
 B) gamma radiation, alpha particle, beta particle
 C) positron, alpha particle, neutron
 D) neutron, positron, alpha particle
- Which list of nuclear emissions is arranged in order from the least penetrating power to the greatest penetrating power?
 A) alpha particle, beta particle, gamma ray
 B) alpha particle, gamma ray, beta particle
 C) gamma ray, beta particle, alpha particle
 D) beta particle, alpha particle, gamma ray
- Base your answer to the following question on Given the nuclear equation:
 $^1_1\text{H} + X \rightarrow ^6_3\text{Li} + ^4_2\text{He}$
 The particle represented by X is
 A) ^9_4Li B) ^9_4Be C) $^{10}_5\text{Be}$ D) $^{10}_6\text{C}$
- Positrons are spontaneously emitted from the nuclei of
 A) potassium-37 B) radium-226
 C) nitrogen-16 D) thorium-232
- What is the decay mode of ^{37}K ?
 A) β^- B) β^+ C) γ D) α
- Which equation represents positron decay?
 A) $^{87}_{37}\text{Rb} \rightarrow ^0_{-1}\text{e} + ^{87}_{38}\text{Sr}$
 B) $^{237}_{92}\text{U} \rightarrow ^{233}_{90}\text{Th} + ^4_2\text{He}$
 C) $^{27}_{13}\text{Al} + ^4_2\text{He} \rightarrow ^{30}_{15}\text{P} + ^1_0\text{n}$
 D) $^{11}_6\text{C} \rightarrow ^0_{+1}\text{e} + ^{11}_5\text{B}$
- As ^{14}C decays to ^{14}N , the number of protons in the nucleus
 A) decreases B) increases
 C) remains the same
- Atoms of I-131 spontaneously decay when the
 A) stable nuclei emit alpha particles
 B) stable nuclei emit beta particles
 C) unstable nuclei emit alpha particles
 D) unstable nuclei emit beta particles
- Which statement best describes gamma radiation?
 A) It has a mass of 1 and a charge of 1.
 B) It has a mass of 0 and a charge of -1.
 C) It has a mass of 0 and a charge of 0.
 D) It has a mass of 4 and a charge of +2.
- Which nuclear decay emission consists of energy, only?
 A) alpha particle B) beta particle
 C) gamma radiation D) positron
- Which kind of nuclear radiation has high energy and no mass?
 A) alpha B) beta
 C) gamma D) neutron
- The half-life of a radioactive isotope is 20.0 minutes. What is the total amount of a 1.00-gram sample of this isotope remaining after 1.00 hour?
 A) 0.500 g B) 0.333 g
 C) 0.250 g D) 0.125 g

Half-life	Amount remaining
0	1.00 g
20.0 min	0.500 g
40.0 min	0.250 g
60.0 min	0.125 g
- A radioactive element has a half-life of 2 days. Which fraction represents the amount of an original sample of this element remaining after 6 days?
 A) $\frac{1}{8}$
 B) $\frac{1}{2}$
 C) $\frac{1}{3}$
 D) $\frac{1}{4}$

0	1/1
2	1/2
4	1/4
6	

Review Nuclear Chemistry Section 2 Answers

Gary Wulfsberg



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McGraw-Hill's SAT Subject Test: Chemistry, 2ed Thomas Evangelist, 2009-02-01 We want to help you score high on the SAT Chemistry test. We've put all of our proven expertise into McGraw Hill's SAT Subject Test Chemistry to make sure you're fully prepared for this difficult exam. With this book you'll get essential skill building techniques and strategies created by a leading high school chemistry teacher. You'll also get 5 full length practice tests, hundreds of sample questions and all the facts about the current exam. With McGraw Hill's SAT Subject Test Chemistry we'll guide you step by step through your preparation program and give you the tools you need to succeed. 5 full length sample tests with complete explanations for every question. 40 top test items to remember on exam day. A step by step review of all topics covered on the exam. Teacher recommended tips and strategies to help you raise your score.

Kaplan SAT Subject Test Chemistry 2015-2016 Kaplan Test Prep, 2015-03-03 Essential strategies, practice and review to ace the SAT Subject Test Chemistry. Getting into a top college has never been more difficult. Students need to distinguish themselves from the crowd and scoring well on a SAT Subject Test gives students a competitive edge. Kaplan's SAT Subject Test Chemistry is the most up to date guide on the market with complete coverage of both the content review and strategies students need for success on test day. Kaplan's SAT Subject Test Chemistry features: A full length diagnostic test. Full length practice tests. Focused chapter summaries, highlights and quizzes. Detailed answer explanations. Proven score raising strategies. End of chapter quizzes. Kaplan is serious about raising students' scores; we guarantee students will get a higher score.

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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001

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