

Class 45

SECTION 3:

NOTE: Answer each multiple-choice question in the space provided.

1. Which solution has the greatest set of ions in solution? Choose, or write the name in the following-instances:

- (a) 1.0 M aqueous solution: sodium chloride solution
- (b) 1.0 M aqueous: magnesium chloride solution
- (c) 1.0 M aqueous solution: sodium hydroxide solution
- (d) 1.0 M aqueous solution: magnesium sulfate solution

2. Two sealed flasks, A and B, contain gases at equal volume in the same temperature conditions. Gas A has a molecular weight of 44 g/mol and gas B has a molecular weight of 32 g/mol. Which flask contains more gas? If not, which flask will have the higher pressure? Which pressure will be lower? Explain. (Half of the flask contains gas A, and the other half contains gas B.)

3. Which gas has the highest average molecular weight? Which gas has the highest average molecular weight?

4. A gas is at 1.0 atm pressure at 25°C. If the volume of the gas is doubled, what happens to the pressure?

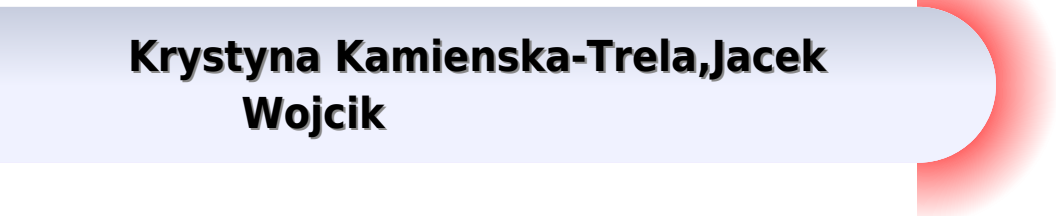
5. If the volume of the gas is doubled at 25°C, what happens to the pressure? If the volume of the gas is doubled, what happens to the pressure? (The volume of the gas is doubled.)

6. If the volume of the gas is doubled at 25°C, what happens to the pressure? (The volume of the gas is doubled.)

7. If the volume of the gas is doubled at 25°C, what happens to the pressure? (The volume of the gas is doubled.)

Review Gases Section 2 Short Answer

**Krystyna Kamienska-Trela, Jacek
Wojcik**



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Review Gases Section 2 Short Answer Book Review: Unveiling the Power of Words

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