

Equilibrium: Checking for Understanding & Practice Problems

Name Bern "key" Period

As a reaction is approaching equilibrium describe how the following change.

1. Reactant concentration.
As the rxn goes to the right the reactant concentration decreases.
2. Products concentration.
As the rxn goes to the right the product concentration increases.
3. Forward reaction rate.
The reactant concentration decreases as the rxn proceeds, therefore there are less reactant collisions causing the forward rate to decrease.
4. Reverse reaction rate.
As the rxn proceeds, the product concentrations ↑ so more product collisions ↑ the rate of the reverse rxn.
5. What is equal at equilibrium?
The rate of the forward rxn and the rate of the reverse rxn.
6. What is constant at equilibrium?
The reactant and product concentrations.

Solve the following problems and be sure to show all of your work.

1. At equilibrium at 100°C, a 2.0L flask contains:

0.075 mol of PCl_5 0.050 mol of H_2O 0.750 mol of HCl 0.500 mol of POCl_3
Calculate the K_{eq} for the reaction:



$$M = \frac{\text{mol}}{\text{L}}$$

$$M_{\text{H}_2\text{O}} = \frac{0.050 \text{ mol}}{2.0 \text{ L}} = 0.025 \text{ M}$$

$$M_{\text{HCl}} = \frac{0.750 \text{ mol}}{2.0 \text{ L}} = 0.38 \text{ M}$$

$$M_{\text{POCl}_3} = \frac{0.500 \text{ mol}}{2.0 \text{ L}} = 0.25 \text{ M}$$

$$K_{eq} = \frac{[\text{HCl}]^2 [\text{POCl}_3]}{[\text{H}_2\text{O}]}$$

$$K_{eq} = \frac{(0.38)^2 (0.25)}{(0.025)} = 1.4$$

$$K_{eq} = 1.4$$

Review Chemical Equilibrium Answers Section 1

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