

**CHAPTER 6 REVIEW***Chemical Bonding***SECTION 4****SHORT ANSWER** Answer the following questions in the space provided.

1. \_\_\_\_\_ In metals, the valence electrons are considered to be
  - (a) attached to particular positive ions.
  - (b) shared by all surrounding atoms.
  - (c) immobile.
  - (d) involved in covalent bonds.
2. \_\_\_\_\_ The fact that metals are malleable and ionic crystals are brittle is best explained in terms of their
  - (a) chemical bonds.
  - (b) London forces.
  - (c) enthalpies of vaporization.
  - (d) polarity.
3. \_\_\_\_\_ As light strikes the surface of a metal, the electrons in the electron sea
  - (a) allow the light to pass through.
  - (b) become attached to particular positive ions.
  - (c) fall to lower energy levels.
  - (d) absorb and re-emit the light.
4. \_\_\_\_\_ Mobile electrons in the metallic bond are responsible for
  - (a) luster.
  - (b) thermal conductivity.
  - (c) electrical conductivity.
  - (d) All of the above.
5. \_\_\_\_\_ In general, the strength of the metallic bond \_\_\_\_\_ moving from left to right on any row of the periodic table.
  - (a) increases
  - (b) decreases
  - (c) remains the same
  - (d) varies
6. \_\_\_\_\_ When a metal is drawn into a wire, the metallic bonds
  - (a) break easily.
  - (b) break with difficulty.
  - (c) do not break.
  - (d) become ionic bonds.
7. Use the concept of electron configurations to explain why the number of valence electrons in metals tends to be less than the number in most nonmetals.  
\_\_\_\_\_  
\_\_\_\_\_  
\_\_\_\_\_  
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# Review Sheet Chemical Bonding Answers

**Sasha Hu**



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