



Restoring Balance

LeChâtelier's Principle and Equilibrium

Introduction

Chemical equilibrium is a unue balancing au. What happens when the balance is disturbed. The purpose of this lab is to observe the effects of concentration and temperature on equilibrium and to visualize how balance can be restored based on LeChitelier's Principle.

Concepts

· Chemical equilibrium

· LeChäselier's principle

Complex-ton residues

Exactlements via enclustrements reactions.

Background

Not, all chemical reactions proceed to completion, giving 100% yield of processes. In fact, most, chemical reactions are reversible, meaning they can go both ways. In the forecast direction, reactions interact to make process while in the reverse direction, the products revers both to reactions. This idea is represented symbolically using disable arrows as shown for the reversible reaction of incline nucleonless and indice into the give principle into the term of the control of the process of the

$$T_i(sc) + T(sc) e^{-s}T_i(sc)$$
 Equation 2

In a closest systems, any reversible reaction will eventually reach a system is signarate behavior between the forward and reverse reactions. A system is said as reach chemical equilibrium when the rate of the forward reaction equals the rate of the reverse reaction. At this point, no further changes will be observed in the ansounts of either the reactions or products. Chemical equilibrium can be further defined, therefore, as the state at which the concentrations of reactions and products remain constant, with time. This does not never the concentrations of reactions and products are called the products are of all the forward and reverse reactions create an equal hallance of opposing rates.

What happens when the balance is disturbed moine to the addition of noise reactants or products or due to changes in the temperature or pressure? LeChitelier's Principle precious how equilibrium can be restored:

"Af an equilibrium system is subjected to a stress, the system will reset in such a way as to reduce the stress."

Any change that is made to a system at equilibrium is considered a stream-this includes adding or removing reagents or changing the temperature. To reduce the stress, one of two things can happen. A reversible reaction can shift in the forward direction and make more products, thus using up some of the reactions. Alternatively, the reaction can shift in the reverse direction and representation are directions and professor the reactions, thus using up some of the products.

The effect of temperature on a system at exhibitions depends on whether a reaction is endothermic (absorbs heat) or exothermic (produces heat). If a reaction is endothermic, best appears on the reactant side in the chemical equation. Increasing the temperature of an endouthermic reaction shifts the equificient in the forest direction to consume some of the excess energy and make more products. The opposite effect is observed for exothermic reactions. In the case of an exothermic reaction, heat appears on the product side in the chemical equation.

There as no present of an exothermic reaction while the equilibrium in the reverse direction.

Experiment Overview

The purpose of this experiment is to investigate the effect of reaction conditions on the reversible formation of subult complex ions. When subult oblives breakly trate (CoCl₂ $^{\circ}$ H₂O) is dissolved in ethyl skudul, three different subult species are present Co^{2} casions Cl anions and vater molecules. There can react to form two different complex ions, $Co(H_{2}O)_{2}^{-1}$, where the collection is surrounded by four chloride torus $Co(H_{2}O)_{2}^{-1}$, in which the needs too is surrounded by four chloride torus.

Pre-Lab Questions

Tourne (I_i) is only sparingly soluble in water (Equation 2). In the
presence of potassium tourier, a source of tourier (I I tions, tourne
reacts to form trimcide (I_i) tons (Equation 3).

$$T_{j}(x) \le T_{j}(xx)$$
 Equation 2

$$I_{g}(sc) + I_{g}(sc) \in I_{g}(sc)$$
 Equation 3

Use DeChâzelier's Principle to explain why the solubility of indire in water increases as the concentration of patassium indice increases.

 Disordising communicate chloricle in water is an encothermic reaction. (Reseation 4). Use LeChitelier's Principle to precion obselver annountary chloricle will be more soluble in hot or cold water.

$$NH_{j}Cl(s) + heat_{j} ? NH_{j}(sc) + Cl(sc)$$
 Equation 4

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Clifford Lane

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